

Mobility and fluxes of trace elements and nutrients at the sediment–water interface of a lagoon under contrasting water column oxygenation conditions

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Abstract:

The early diagenesis of the major carrier phases (Fe and Mn minerals), trace elements (As, Co, Cr, Hg, MeHg, Ni) and nutrients (ΣNO_3 , NH_4^+ , ΣPO_4) and their exchange at the sediment water/interface were studied in the Berre Lagoon, a Mediterranean lagoon in France, at one site under two contrasting oxygenation conditions (strictly anoxic and slightly oxic) and at an adjacent site with perennially well-oxygenated water. From the concentration profiles of the primary biogeochemical constituents and trace elements of the pore and bottom waters, as well as the total and reactive particulate phases, we were able to locate and identify the diagenetic reactions controlling the mobility of trace elements in the sediments and quantify their rates by coupling one-dimensional steady-state transport-reaction modelling and thermodynamic speciation calculations.

Under oxic conditions and in the absence of benthic organisms, the main redox reactions were well identified vertically in the surface sediments and followed the theoretical sequence of oxidant consumption: $\text{O}_2 > \Sigma\text{NO}_3/\text{MnO}_2 > \text{Fe}(\text{OH})_3 > \text{SO}_4^{2-}$. However, under anoxic conditions, only MnO_2 , $\text{Fe}(\text{OH})_3$ and SO_4^{2-} reduction were present, and they all occurred at the interface. We identified the main biogeochemical controls on the mobility of As, Cr, Hg, MeHg and Ni in the surface sediments as the adsorption/desorption and/or coprecipitation/codissolution with Fe oxy-hydroxides. In contrast, Co mobility was primarily controlled by its reactivity towards Mn oxy-hydroxides. In sulphidic sediments, As, Hg and MeHg were sequestered along with Fe sulphides, whereas Co and Ni precipitated directly as metallic sulphides and Cr mobility was enhanced by complexation with dissolved organic ligands. The fluxes of trace elements at the sediment-water interface are essentially dependent on the localisation of their remobilisation and immobilisation reactions under the interface, which in turn is

governed by the benthic water oxygenation conditions and kinetic competition among those reaction and diffusion processes. Under oxic conditions, the precipitation of Fe or Mn oxy-hydroxides in the surface sediments constitutes the most efficient mechanism to sequester most of the trace elements studied, thus preventing their diffusion to the water column. Under anoxic conditions the export of trace elements to the water column is dependent on the kinetic competition among the remobilisation during the reductive dissolution of Fe and/or Mn oxy-hydroxides, diffusion and immobilisation with sulphides. We also show that benthic organisms in the perennially oxygenated site have a clear impact on this general pattern. Based on our extensive dataset and geochemical modelling, we predict that the planned reoxygenation of the entire lagoon basin, if complete, will most likely limit or reduce the export of the trace elements from the sediments to the water column and therefore limit the impact of the contaminated sediment.

Highlights

► Main constituents and trace elements concentration profiles in contaminated sediments. ► Contrasting oxygen concentrations in the overlying water column. ► Combination of thermodynamic calculations and inverse transport-reactions modelling. ► Identification and quantification of diagenetic reactions controlling trace elements mobility. ► Assessment of trace elements and nutrients fluxes at the sediment–water interface.

1. Introduction

During the last century, most coastal areas were strongly impacted by the development of anthropogenic activities on their shores and the subsequent inputs of contaminants or nutrients. Although these releases were significantly reduced during recent decades because of regulations, contaminants accumulated in the sediment over time. These sediments now constitute a potential source of contaminants for the water column that could alter the water quality and threaten aquatic organisms. Understanding the processes controlling the dynamics of contaminants in the sediment, improving our knowledge of the environmental risk induced by contaminated sediments and appropriately orientating the action of politics and managers, is thus a challenge for the next decades (Luoma and Rainbow, 2008).

Trace metals or metalloids (referred to trace elements hereafter) are among the most ubiquitous contaminants. Their remobilisation in sediments depends on the prevailing physicochemical conditions in both the sediment and water column, and on the biogeochemical processes occurring during the early diagenesis of sediment: biogeochemical reactions (e.g., sorption, (co)precipitation, dissolution) and physical (e.g., diffusion, advection) or biological (e.g., bioirrigation, bioturbation) transport mechanisms. Biogeochemical reactions, primarily controlled by microbial mineralization of organic matter, can increase the concentration of trace elements in porewater, and transport mechanisms can induce their transfer to the water column. Trace element fluxes at the sediment-water interface (SWI) have been observed in coastal areas, but their direction and intensity vary among sites and elements (e.g., Point et al., 2007). Among the factors influencing the intensity of the fluxes at the SWI, the oxygen concentrations in the water column was often identified as the most important one because it influences all of the biogeochemical characteristics of sediments (Middelburg and Levin, 2009). A low oxygen concentration in the water column has been found to enhance the flux of Fe, Mn, As, Co, Hg and MeHg from sediment to the water column (e.g. Sundby et al., 1986; Emili et al., 2011), whereas a high oxygen concentration can enhance the flux of Cu from the water column to the sediment (Riedel et al., 1997). The oxygenation of the water column also strongly influences the fluxes of other species important for environmental quality such as ΣPO_4 , NO_3^- , NH_4^+ and H_2S . For example, under hypoxic or anoxic conditions, fluxes of ΣPO_4 , H_2S and NH_4^+ to the water column and fluxes of NO_3^- to the sediment are usually observed whereas in oxic conditions these fluxes tend to stop or reverse (Sundby et al. 1986; Middelburg and Levin, 2009; Emili et al., 2011).

The influence of the oxygen concentration of the water column on reactive processes in the sediment and fluxes at the SWI is still poorly understood, especially in areas where the oxygen concentration of the water column oscillates seasonally (Morse and Eldridge, 2007). Furthermore, the increasing number of coastal areas subject to seasonal or permanent hypoxic episodes in recent decades (Diaz and Rosenberg, 2008) heightens the importance of understanding the diagenetic processes and fluxes at the SWI under hypoxic conditions.

The objective of this work is to study the main diagenetic reactions controlling the mobility of trace elements and nutrients in sediments and their exchange at the SWI under varying water column oxygen conditions using a method that has been successfully used in many different aquatic environments (Gallon et al., 2004; Merrit and Amirbahman, 2007; Couture et al., 2010); it consists of two complementary approaches: 1) one-dimensional steady-state inverse modelling of transport-reaction processes in porewater and 2) thermodynamic calculations (speciation and saturation indices). This method is applied here on the most extensive database of its kind, to our knowledge, which, for the first time, is applied to several trace metals (As, Co, Cr, Hg, MeHg, Ni), nutrients and main chemical species collected at one site under two contrasting water column oxygen conditions and at a second site with oxic condition and macrobenthic organisms.

The study area is the Berre Lagoon (Southeast, France), a heavily industrialised coastal lagoon that received significant inputs of contaminants during the previous decades. These inputs have been regulated since the 1970's, but the concentrations of contaminant in surface sediments remain high (Rigaud et al., 2011). This lagoon also receives freshwater inputs from a hydroelectric power plant, causing a strong haline stratification of the water column and seasonal anoxic conditions in the deepest part of the lagoon. In the future, this freshwater release may be reduced or stopped, that may allow the destratification of the water column and the disappearance of these anoxic conditions. Such a change will certainly have a strong influence on the trace elements and nutrient exchanges at the SWI, which must be evaluated.

In this context, the objectives of this paper are (i) to identify the main biogeochemical reactions controlling trace element mobility in the sediments of the Berre Lagoon under different water column oxygen concentrations, (ii) to quantify trace element and nutrient exchanges at the SWI and (iii) to predict the influence on these processes and fluxes of the reoxygenation of the water column in the deepest part of the lagoon.

2. Material and methods

2.1. Sampling

Two sampling sites were examined in this work: site V located in the Vaïne basin, which had a generally well-oxygenated water column, and site CB located in the southern part of the Central basin, where anoxic/hypoxic events occurred in the benthic water (Fig. 1, Table 1). Benthic organisms are permanently absent at site CB due to the occurrence of anoxic events. Site V was sampled in June 2010 during oxic conditions (V-O, $[O_2] = 4.6 \text{ mg.L}^{-1}$; i.e. 60% of saturation). Site CB was sampled on two occasions, in summer 2009 during anoxic conditions ($[O_2] < 0.2 \text{ mg.L}^{-1}$; < 3% of saturation) and in summer 2010 during slight oxic conditions ($[O_2] = 2.9 \text{ mg.L}^{-1}$; 38% of saturation), hereafter referred to as CB-A and CB-O, respectively.

Three sediment cores were collected by divers using polycarbonate tubes (inner diameter 9.4 cm), placed in refrigerated container, transported to the laboratory and processed within two hours. Eh were measured by inserting probes directly into the sediment through holes in one specific predrilled tube; these holes were covered with tape during sampling. One core,

dedicated to solid phase characterisation, was sliced at 0.5 cm intervals (from 0 to 1 cm), 1 cm intervals (from 1 to 5 cm) and 2 cm intervals (from 5 cm to the end, up to 30 cm) and the slices were placed in polyethylene tubes, freeze-dried and sieved through 63 μm sieve to remove shells and broken shells. In order to minimize the oxidation of the solid phase, sediment cores were sliced in a glove box under N_2 atmosphere and the solid phase was kept in closed vials until the freeze-drying step. One core was collected for O_2 profiles measurements by microelectrodes within two hours after collection.

Porewater and overlying waters were collected using dialysis samplers (peepers) consisting of two columns of 50 horizontal cells with 1 cm vertical resolution drilled in a methacrylate casing. Prior to deployment, the peepers were immersed in a 10% HNO_3 solution for three days, rinsed and filled with ultrapure water, covered with a 0.45 μm polycarbonate filter membrane and then stored for three weeks in a water-saturated N_2 atmosphere to ensure complete removal of oxygen. The peepers were inserted into the sediment by divers and left in place for at least three weeks. Upon retrieval, the water in each cell was collected and filtered through a 0.2 μm cellulose acetate filter. This step was carried out on field within one hour after peeper retrieval. During the 2010 field campaign (sites CB-O and V-O), two peepers were deployed, and the water samples were aliquoted for chemical analysis. During the summer 2009 field campaign (site CB-A), four peepers were deployed, and the sub-sampling plan was slightly modified for the additional analysis of three distinct profiles of Hg and MeHg concentrations.

2.2. Analysis

2.2.1. Solid Phase Properties

Sediment grain size was measured using laser granulometry (LS13320, Beckman coulter) following the elimination of organic matter through combustion in an oven at 450°C for 4h and sample dispersion into defloculant ($(\text{NaPO}_3)_6$, 3‰). Sediment water content was determined from mass differences between fresh and freeze-dried sediments and porosity was assessed following correction for salt content and shells (grain size fraction >2 mm) and assuming a sediment density of 2.65 $\text{g}\cdot\text{cm}^{-3}$. The main mineralogical composition of the sediment was obtained using X-ray diffraction (X'Pert PRO, Panalytical).

An aliquot of freeze-dried and <63 μm sieved sediment was completely digested in a Teflon bomb in a microwave oven (Mars5, SEM) with an ultrapure HF, HCl and HNO_3 mixture. Following digestion, the acids were evaporated to near-dryness and the residues were recovered in HNO_3 2.5% and analysed for major (Ca, Mg, S, P, Al, Fe and Mn) and trace (As, Co, Cr, Ni) elements using ICP-AES (Ultima-C, Jobin Yvon, Horiba) and ICP-MS (Thermo Scientific X Series II), respectively.

The most reactive fraction of the major and trace elements was determined using ascorbate extraction (reductive dissolution) on freeze-dried sediments (Kostka and Luther; 1994). Extractions were carried out at room temperature (24°C) on 100 mg of freeze-dried sediment in 10 mL of N_2 -degassed ascorbate reagent (pH 8) for 24 h under continuous shaking. The ascorbate extracts were analysed using ICP-AES or ICP-MS.

Elemental analyses in solutions were performed using calibration solutions prepared with the same matrix as the samples (HNO_3 2.5% or ascorbate). Blanks were always below 5% of the samples concentrations. Average analytical accuracies, estimated for the total sediment digestion of reference sediments SGR-1, MAG-1 (USGS) and STSD-3 (CCRMP) and MAG-1 (USGS), were always within 5% for Ca, Mg, K, S, P, Al, Fe, Mn, As and Ni and approximately 10 % for Co and Cr.

Total Hg (Hg_{tot}) and MeHg concentrations were obtained for freeze-dried sediments using, respectively, a cold vapour atomic absorption spectrophotometer (AMA 254, Altec) and GC-

ICP-MS (GC-Focus with X-series, Thermo Electron) following the protocol described by Abi-Ghanem et al., (2011). The analytical accuracies were obtained using reference sediments, MESS-2 (NRCC) for Hg and IAEA-405 (IAEA) for MeHg, and were <3% for total Hg and approximately 5% for MeHg.

Total organic carbon (TOC) and total inorganic carbon (TIC) concentrations were measured on a Leco 125 CS analyser using the protocol described by Etcheber et al., (1999). Total nitrogen (N_{tot}) concentrations were analysed on a CHN Leco. All measurements performed on freeze-dried solid phase samples were corrected for salt content.

2.2.2. Oxygen profiles measurements by microelectrodes

Oxygen profiles measurements by microelectrodes were carried out at in situ temperature and O_2 concentration into the overlying water column was adjusted to in situ conditions by bubbling an air- N_2 mixing into the water column. Profiles were measured at 100 μm depth increments with a Clark-type O_2 microelectrode (OX50, Unisense, Denmark) equipped with an internal reference and a guard cathode. The sensor tip diameter was 50 μm , the 90% response time was < 5 s, and the stirring sensitivity was < 1s. Microelectrodes were connected to a high-sensitivity picoamperemeter (PA2000, Unisense) via an A/D converter and were mounted on a motor controlled micromanipulator (MC-232, Unisense). Linear calibrations were done between the bottom water oxygen content precisely determined by a Winkler titration, and the zero oxygen in the anoxic part of the sediment. Up to 10 steady-state O_2 profiles were randomly performed at each site in the collected sediment core to account for spatial heterogeneities. The position of the sediment-water interface was determined from a break in the oxygen concentration gradient. After profiles measurements, the core was sliced every 0.2 cm over the first cm, and then every 0.5 cm down to 2 cm for porosity measurements. Oxygen profiles were processed using the PRO₂FLUX program software (Deflandre and Duchêne, 2010).

2.2.3. Dissolved Phase Analysis

Water samples were analysed for Na, Ca and Mg using ICP-AES (Ultima-C, Jobin Yvon, Horiba) and for Fe, Mn, As, Co, Cr and Ni using ICP-MS (Agilent 7700x equipped with HMI kit). The accuracies were checked using a seawater reference NASS-4 (NRCC) enriched with trace elements to achieve concentration ranges similar to the samples. The accuracies of the measurements were within 5% for Na, Ca, Mg, Fe, Mn, Co, Ni and ~10% for As and Cr.

Total dissolved Hg (Hg_{tot}) concentrations were determined using cold vapour atomic fluorescence spectrometry (CVFAS, Tekran). The accuracies were confirmed using the reference material ORMS-4 (NRCC) and were approximately 5%. The analysis of dissolved MeHg was performed using atomic fluorescence spectrophotometry following the protocol developed by Tseng et al. (1998).

The pH measurement and the analysis of total alkalinity, and SO_4^{2-} , ΣH_2S and ΣPO_4 concentrations were performed using spectrophotometry (V-650 UV-VIS spectrophotometer, Jasco) following Metzger et al. (2007). Dissolved organic carbon (DOC) concentrations were determined on a TOC-V CHS analyser, Shimadzu. Total inorganic carbon concentrations, tCO_2 (CO_2 , HCO_3^- and CO_3^{2-}), were obtained from the total alkalinity measurement, pH value and the concentrations of all of the species included in the definition of total alkalinity ($B(OH)_4^-$, $H_3SiO_4^-$, $H_2SiO_4^{2-}$, H_3PO_4 , HPO_4^{2-} , PO_4^{3-} , NH_3 , HS^- , HSO_4^- , H^+ , OH^- and DOC), which were obtained by direct analysis, assessed from salinity and/or thermodynamic calculations (see below). NH_4^+ and ΣNO_3 ($NO_3^- + NO_2^-$) concentrations were measured by flow injection analysis (FIA) following the methods developed by Hall and Aller (1992) for NH_4^+ and Anderson (1979) for ΣNO_3 .

2.3. Thermodynamic Calculations

The speciation of dissolved species in interstitial and overlying water was calculated through the consecutive use of two geochemical speciation codes: WHAM-VI (Tipping et al., 1998) and PHREEQC (Parkhurst and Appelo, 1999). This approach benefits from the specific strengths of each code, namely the calculation of the complexation among the trace elements and the dissolved organic ligands (WHAM-VI) and of redox equilibrium and saturation index (SI) calculation (PHREEQC). The method includes 3 successive steps: 1) the speciation of the trace elements with fulvic acids (FA) and humic acids (HA) is calculated using WHAM-VI; 2) the apparent thermodynamic constants (K_{app}) of the complexation of the trace elements with HA and FA are calculated (*electronic annex EA-2*) and integrated into the MINTEQ database (used as a database for PHREEQC); and 3) the speciation of the trace elements and their saturation index with respect to the mineral phases are calculated by the software PHREEQC.

Prior to the calculation, the MINTEQ and WHAM-VI default databases were completed from the thermodynamic constants given in *electronic annexes EA-3 and EA-4*. The input parameters used for WHAM-VI were the pH and total dissolved concentrations in HA, FA, Cl, Na, Mg, K, SO_4 , tCO_2 , PO_4 , HS, NH_4 , NO_3 , Fe, Mn, As, Co, Cr, Ni, Hg and MeHg. The concentrations of FA and HA were estimated from a ratio organic matter/DOC of 2/1 (Buffle, 1988) and an FA/HA ratio of 9/1 (Malcolm, 1985). The concentrations of Cl^- and K^+ were estimated from measurements of Na and Dittmar's principle. For PHREEQC calculations, the redox potential Eh was added as input parameter and redox couples in the database were de-coupled to allow redox disequilibrium to occur.

3. Résultats

3.1. Solid Phase

3.1.1. Total Sediment Characteristics

The sediments of the three cores consist of 51-81% silt (2-63 μm), 13-28% clay (<2 μm) and <20% of fine sands (63-250 μm). The grain size fraction > 2 mm (mainly shells or shell fragments) is absent in CB-O and CB-A and is <15% dry weight in V-O. The fraction on which the analyses were performed ($f < 63 \mu m$) is constant along the profiles in the three cores (silt: $81 \pm 3\%$ at sites CB-A and CB-O and $85 \pm 1\%$ at site V). The mineral composition is similar among sites and along the profile with the predominance of calcite and aragonite, quartz, pyrite (FeS_2), muscovite ($KAl_2[(OH,F)_2AlSi_3O_{10}]$) and small proportions of dolomite ($CaMg(CO_3)_2$).

The major element concentrations of sediments are relatively homogeneous with depth and similar in the three cores: TIC ($4.8 \pm 0.4\%$), Ca ($150 \pm 16 \text{ mg.g}^{-1}$), Al ($39 \pm 9 \text{ mg.g}^{-1}$), Fe ($21 \pm 2 \text{ mg.g}^{-1}$), S ($13 \pm 2 \text{ mg.g}^{-1}$), Mg ($9 \pm 1 \text{ mg.g}^{-1}$) and Mn ($0.37 \pm 0.07 \text{ mg.g}^{-1}$) (examples in Fig. 2a-f). However, depth profiles of TOC, N_{tot} and P_{tot} show concentration maxima at the surface. Significantly higher TOC and N_{tot} concentrations are observed at V-O (Fig. 2a-b).

Trace element concentrations vary greatly with depth, and most of trace elements are at high concentrations a few centimetres below the SWI: the highest As and Hg_{tot} concentrations occurs at approximately -15 cm at sites CB-O and CB-A and at 20-25 cm at site V-O, and the highest Co, Cr and Ni concentrations occur at -10 cm at site V-O (Fig. 2g-j). MeHg, only analysed at CB-A, is at the highest concentrations at the SWI.

A comparison of the trace elements concentrations with the natural geochemical background for the area (Rigaud et al., 2011) shows that Hg and As at the three sites and Cr at site V-O are strongly enriched in the sediment (e.g., up to 14 times the natural geochemical

background for Hg at sites CB-A and CB-O, Fig. 2j). These enrichments are due to the significant releases into the lagoon in the past, whereas their decrease in the upper part of the sediment cores reflects the establishment of industrial regulations (Rigaud et al., 2011).

3.1.2. Reactive Fraction of the Solid Phase

Ascorbate extractions give the primary information on the main oxidised phases, such as reactive Fe or Mn oxy-hydroxides (referred hereafter to $\text{Fe}(\text{OH})_3$ and MnO_2 for simplicity) that are both oxidants for the mineralisation of organic matter and bearing phases for trace elements (e.g., Burdige, 1993). However, the contribution of other phases, such as FeS oxidised during sample handling is not excluded (Anshutz et al., 1998). In the 3 cores, the concentrations Fe extracted by ascorbate (Fe_{ASC} , Fig. 3a) are the highest at the SWI and decline sharply toward -3 cm. They remain relatively low and constant at deeper levels in the CB-O and CB-A cores. However, in V-O core, the Fe_{ASC} concentrations are two times greater than in the CB cores and decrease slightly with depth until -12 cm.

Surface sediment enrichments in reactive Mn extracted by ascorbate (Mn_{ASC} , Fig. 3b) are found in the CB-A and V-O cores but not in CB-O. As observed for Fe_{ASC} , Mn_{ASC} is also clearly more enriched at V-O. The concentrations of Fe_{ASC} and Mn_{ASC} and their relative proportion with respect to total Fe and Mn in surface sediment are summarised in Table 2.

The concentrations of As_{ASC} , Co_{ASC} , Cr_{ASC} and Ni_{ASC} broadly follow the pattern of their total concentrations in the sediments, with maximum values encountered where the highest total concentrations occurred (Fig. 3c-f). The proportions of the total elements extracted are large for As (up to 70%) but smaller for Co (<35%), Cr and Ni (<20%). However As_{ASC} , Co_{ASC} , Cr_{ASC} and Ni_{ASC} are also highly enriched at the SWI at V-O. Although lower, significant enrichments are also observed at the SWI for As_{ASC} , Ni_{ASC} at sites CB-A and CB-O and for Co_{ASC} only at site CB-A. Interestingly, the concentrations of Co_{ASC} are really low in the first centimetre of surface sediment at site CB-O.

3.2. Porewater

3.2.1. Physicochemical Characteristics

Salinity (electronic annex AE-1), estimated from the Dittmar's principle and measured Na concentrations (Fig. 5a), slightly decreases from the overlying water column to the sediment (21 to 18 at V-O, 32 to 30 at CB-O and 36 to 30 at CB-A). Values of pH are relatively constant over the entire profile at each site: 7.53 ± 0.03 at CB-A, 7.91 ± 0.09 at CB-O and 8.04 ± 0.10 at V-O (electronic annex EA-1). Redox potential values were >350 mV/NHE in the overlying water when oxygen was present (CB-O and V-O) and approximately 50 mV during anoxic conditions (CB-A) (Table 1); they dropped very strongly at the SWI and remained constant at approximately -200/-220 mV at sites CB-O and V-O below -5 cm (electronic annex AE-1). At site CB-A where Eh measurements were not carried out, Eh was fixed to -200 mV.

3.2.2. Major Dissolved Species

Oxygen penetration depth into the sediment (Fig. 4) varies from 0.45 mm at site CB-O to 2.0 mm at site V-O where the occurrence of oxygen peaks are observed up to 4.6 mm depth. Concentrations of Ca, Mg and SO_4^{2-} decrease in the sediment compared with the overlying water. This decrease occurred directly at the SWI at CB-A and at -5 cm at CB-O and V-O (Fig. 5b,c,e), depths that coincide with the appearance and increasing concentrations of $\Sigma\text{H}_2\text{S}$. At site CB-A, $\Sigma\text{H}_2\text{S}$ is also present in the water column (~100 μM).

The concentrations of ΣCO_2 , ΣPO_4 , NH_4^+ and DOC are constant and low in the water column, and increase sharply in the sediment from the SWI (at CB-A and V-O) or from -5 cm (at CB-O and V-O for ΣPO_4) (Fig. 5d,i,j,l). Note that the porewater concentrations of DOC and ΣNH_4^+ at V-O are significantly higher than those at the other sites. The ΣNO_3 concentrations have maxima concentrations in the water column at CB-O and V-O and approach the detection limit in the sediments at the depth where NH_4^+ appears. At CB-A, the ΣNO_3 concentrations are always below the detection limit.

The Fe concentrations at CB-A are highest in the anoxic water column (0.8-1.45 μM , maximum values just above the SWI) and decrease sharply in the sediment down to the limit of detection (Fig. 5g). At CB-O and V-O, the opposite is observed, with concentrations in the oxygenated water column near the limit of detection and relatively high concentrations in the porewater, which are characterised by two peaks at 4 μM at -7 cm and -2 cm (V-O) and a broad peak at 1 μM between -9 and -11 cm (CB-O).

The profile of the Mn concentrations at site CB-O features a high peak reaching 30 μM in the first 4 cm below the SWI and relatively high concentrations in the water column (5-10 μM) (Fig. 5h). At this site, the Mn peak is located above the Fe peak. The Mn profile at CB-A is very similar but with lower concentrations (peak at 2.4 μM in the first 3 cm below the SWI and concentration in the water column of 1.5 μM), but here, the peak of Mn is located below that of Fe. At V-O site, the Mn concentrations approach the limit of detection except around the SWI, where two peaks at +2 cm (0.8 μM) and -7 cm (8.5 μM) are present. For the three sites, the concentrations of Mn decreased with increasing depth to values close to the limit of quantification below -20 cm.

3.2.3. Dissolved Trace Elements

At sites CB-A and CB-O, As and Ni concentrations are highest in the overlying water (As = 20-50 nM, Ni = 10-25 nM) and decrease under the SWI to 5-30 nM and 4 nM, respectively (Fig. 5m,p). However, As exhibits a reverse trend at site V-O with minimum concentrations in the overlying water (10-20 nM) and an increase in the porewater (20-50 nM). At this site, Ni concentrations are rather stable over the entire profile, at approximately 15-25 nM.

The profiles of Co at CB-O and V-O are very similar to those of Mn, with the highest concentrations in the first 4 centimetres below the SWI (up to 50 nM) at CB-O and at +2 cm (10 nM) and -6 cm (8.5 nM) at V-O (Fig. 5n). However, the Co profile clearly differs from that of Mn at CB-A with minimum Co concentrations at the SWI (1 nM) and an increase with depth (up to 12 nM).

The concentrations profiles of Cr differ greatly among the three sites (Fig. 5o). Cr concentrations are low (<35 nM) throughout the profile at CB-O, more variables at CB-A with maxima at the SWI (90 nM) and -12 cm (70 nM), and increase with depth at site V-O with a peak (120 nM) at -10 cm.

The three concentrations profiles of total Hg (Hg_{tot}) and MeHg (site CB-A) are very similar to each other. They show maxima in the overlying water (10-20 pM and 6-11 pM, respectively), strong decreases below the SWI and reach the limit of quantification below -5 cm (Fig. 5q,r). MeHg dominates the speciation of Hg with average proportions of $73 \pm 22\%$ (range 32-108%) in the porewater.

4. Discussion

4.1. Identification of the Main Biogeochemical Reactions

The main early diagenetic reactions as well as those involving the trace elements considered in this work (Table 3) correspond to (i) the primary redox reactions (R1-R6) for the mineralisation of organic matter (OM), (ii) the secondary redox reactions (R7-R21) involving the products from the primary redox reactions and the remaining oxidants, (iii) the non-redox precipitation reactions (R22-R26) and (iv) the reactions controlling the trace element dynamics at the solid/liquid interface, and thus their mobility in the sediment (R27-R32) (e.g., Wang and Van Capellen, 1996).

The approach used here to identify and quantify the biogeochemical reactions controlling the remobilisation or immobilisation of trace elements and nutrients in sediments and their fluxes at the SWI is based on those recently used to determinate the dynamic of Pb (Gallon et al., 2004), Hg and MeHg (e.g., Merrit and Amirbahman, 2007) and As (Couture et al., 2010) in different aquatic environments.

First, the depth intervals where remobilisation or immobilisation reactions occurred in the sediments and the quantification of their net rates are obtained by inverse transport-reaction modelling of the concentration profiles of the species in the porewater. The mass conservation equation in one dimension of a dissolved species as proposed by Boudreau (1997) is

$$\left(\frac{\partial\phi C}{\partial t}\right) = \frac{\partial}{\partial x}\left(\phi(D_S + D_B)\frac{\partial C}{\partial x}\right) - v\frac{\partial C}{\partial x} + \phi\alpha(C_\infty - C) + R_{net} \quad (\text{Equation 1})$$

where C is the concentration (μM) of a dissolved species at time t (s) and depth x (cm, set at SWI and negative below), ϕ is the porosity, D_S ($\text{cm}^2.\text{s}^{-1}$) is the molecular diffusion coefficient in the porewater, D_B ($\text{cm}^2.\text{s}^{-1}$) is the biodiffusion coefficient associated to the sediment reworking process, v is the velocity of the advective flow in porewater (cm s^{-1}), α (s^{-1}) is the bioirrigation rate coefficient, C_∞ (μM) is the concentrations in the water column, and R_{net} ($\mu\text{mol.L}^{-1}.\text{s}^{-1}$) is the net rate of reactions, corresponding to the balance between the reactions of production or consumption of the species in the porewater.

Several hypotheses have been proposed to simplify this equation. First, we assume that the dissolved profiles are at a quasi-stationary state ($\partial C/\partial t \approx 0$). Although this assumption is not completely valid in environments subject to changing conditions, it has been useful to explore the parameters responsible for the immobilisation or remobilisation of trace elements (e.g., Couture et al., 2010). Second, porewater advection due to sediment compaction (Peclet number, $\text{Pe} < 0.1$; Boudreau, 1997) is neglected ($v = 0$). Third, the recurrent anoxic conditions in the water column at site CB totally prevent the colonization of the sediment by benthic organisms (Fig. 1), and consequently, sediment reworking and bioirrigation are absent in CB-A and CB-O (D_B and $\alpha = 0$). Because the benthic faunal density in the Vaïne basin is low ($< 500 \text{ ind.m}^{-2}$, GIPREB, com. pers.), biodiffusion and bioirrigation were also neglected at site V-O. This simplification has a negligible effect for the parameter D_B that is generally at least an order of magnitude lower than D_S (Schulz, 2006). However, not considering possible transport by bioirrigation (i.e.; assuming $\alpha = 0$) could lead to an underestimation of the fluxes and reaction rates at site V-O, but such simplification doesn't impact the general discussion of process occurring in the sediment.

From these assumptions, Equation 1 is rewritten to express the net rates of reactions only as a function of the diffusion of dissolved species in the porewater (Equation 2).

$$R_{net} = \frac{\partial}{\partial x}\left(\phi D_S \frac{\partial C}{\partial x}\right) \quad (\text{Equation 2})$$

Equation 2 was solved numerically using the code PROFILE (Berg et al., 1998), which yields positive R_{net} values for species production in the porewater and negative for species consumption in the porewater. The values of D_s are obtained from the coefficient of molecular diffusion in free water D^0 corrected for salinity, temperature and sediment tortuosity (Boudreau, 1997). For complexed species for which there are no D^0 data, the values for the free ion are used. For Cr, which is predominantly complexed to the OM, the D^0 of fulvic acid is used (Gallon et al., 2004).

Finally, the reactions involved in the remobilisation or immobilisation processes are identified assuming that if two species (dissolved or particulate) involved in a reaction in Table 3 are present in measurable concentrations at the same depth, the reaction is considered. In the specific case of an area where inverse modelling suggests that the consumption of a dissolved species ($R_{\text{net}} < 0$) is linked to solute transfer from the porewater to the particle, we take into account two processes: i) precipitation reaction is considered if thermodynamic calculations predict that a dissolved species reach saturation with respect to a mineral phase, ii) (ad)sorption and/or coprecipitation reactions are considered if thermodynamic calculations do not reveal any favourable precipitation reactions.

4.2. Main Early Diagenetic Reactions

4.2.1. Evidence for OM mineralisation

At the three sites, the decrease with depth of particulate TOC (Fig. 2a), in parallel with the increase of ΣCO_2 , NH_4^+ , ΣPO_4 and DOC porewater concentrations (Fig. 5d,i,j,l) directly reflect the mineralisation of OM during burial (R1 to R6, Table 3) and resulting redox reactions.

4.2.2. Primary and secondary redox and precipitation reactions

Measured and modelled porewater profiles of ΣNO_3 , Fe, Mn, SO_4^{2-} , $\Sigma\text{H}_2\text{S}$ and the trace elements are presented in Fig. 6A, Fig. 7A and Fig. 8A. A schematic representation of the vertical distribution of the major and trace dissolved species in the porewater is also proposed for each site (Fig. 6B, Fig. 7B and Fig. 8B). The schematic representation reports (i) the relative concentrations, (ii) the areas of production and consumption, (iii) the reactions assumed to be involved (the numbers of reactions reported in Table 3), (iv) the net rates of reactions integrated over the depth interval (R_{net} in $\text{mol}\cdot\text{cm}^{-2}\cdot\text{s}^{-1}$) and (v) the significant diffusive fluxes at the SWI ($\text{mol}\cdot\text{cm}^{-2}\cdot\text{s}^{-1}$) and their direction. The main identified biogeochemical reactions and their locations in the sediment are summarised in Table 4.

At site CB-O (Fig. 6), the reactions involved in the OM mineralisation are well resolved vertically and follow the theoretical sequence proposed by Froelich et al. (1979): consumption of O_2 (within the first 0.45 mm), ΣNO_3 reduction (down to -5.7 cm depth), MnO_2 reduction (0 to -2.4 cm depth), $\text{Fe}(\text{OH})_3$ reduction (-9.5 to -11.9 cm depth) and SO_4^{2-} reduction (-14.3 to -28.6 cm depth). The intense MnO_2 reduction below the SWI is consistent with the low reactive Mn_{ASC} concentrations in the surface sediment (Table 2; Fig. 3b). The Mn produced in the porewater diffuses deeper into the sediment and is completely consumed between -2.4 and -7.2 cm. This consumption is not related to the precipitation of the usual Mn minerals, which are under-saturated (rhodochrosite, MnCO_3 , $\text{SI} < -1$; MnS or haurite; $\text{SI} < -2$), but could rather be associated with the precipitation of another form of Mn carbonate or coprecipitation with carbonates (R31) as previously reported (e.g., Pingitore et al. 1988; Deflandre et al., 2002). The consumption of Fe between 0 and -9.5 cm is associated with the precipitation of $\text{Fe}(\text{OH})_3$ (R14, R17), which contributes to the enrichment of Fe_{ASC} at the SWI (Table 2; Fig. 3a). Such precipitation is furthermore supported by the saturation or oversaturation of Fe with respect to several Fe oxy-hydroxydes (hematite, magnetite, lepidocrocite, $\text{SI} \geq 0$). More deeply in the porewater, Fe is rapidly consumed through Fe

sulphide precipitation (R25). This is consistent with i) the saturation of the porewater with respect to pyrrhotite and troilite ($0 < SI < 0.7$); ii) the occurrence of pyrite, which requires for its formation an intermediate step involving Fe monosulphides (R25 + R26; Rickard and Morse, 2005); and iii) the significantly lower depth integrated production rate of $\Sigma\text{H}_2\text{S}$ ($R_{\text{net}} = 2.90 \cdot 10^{-13} \text{ mol.cm}^{-2} \cdot \text{s}^{-1}$) in comparison with the depth integrated consumption rate of SO_4^{2-} ($R_{\text{net}} = -2.33 \cdot 10^{-12} \text{ mol.cm}^{-2} \cdot \text{s}^{-1}$), indicating a loss of S in the porewater.

At the same site under anoxic conditions (CB-A, Fig. 7), the only available oxidants for the mineralisation of OM are MnO_2 , $\text{Fe}(\text{OH})_3$ and SO_4^{2-} , whereas O_2 and ΣNO_3 are depleted above the SWI. In contrast with the oxic conditions, the reductions of these oxidants all occur at the SWI with the production of Mn (R3, R16 and R18) between 0 to -3.7 cm depth, the production of Fe (R4, R19, R20) being only noticeable at the SWI (a single data point that was not captured by PROFILE) and the consumption of SO_4^{2-} (R5) between 0 and -5.0 cm depth (Table 4). The modelled porewater Fe profile (Fig. 7A) displays a very high consumption rate between 0 and -2 cm in depth, which is attributed to the precipitation of FeS (R25). The occurrence of this precipitation in surface sediment strongly suggest i) that the reduction of $\text{Fe}(\text{OH})_3$ and the precipitation of FeS both occur at the same depth and ii) that the kinetics of FeS precipitation dominates the reduction $\text{Fe}(\text{OH})_3$, hiding the production of Fe in porewater at the SWI. This explains the location of the Fe peak above the Mn peak.

At site V-O, where the overlying water is oxygenated and macrofauna are observed in the sediment, the identification of the reactions is more complex. Two overlapped diagenetic sequences can be observed (Fig. 8). The first sequence spans across the SWI and involves the reduction of ΣNO_3 (R2; between 0 and +5 cm) and MnO_2 (R3, R16; between +1 and +3 cm) in the overlying water (not modelled by PROFILE) and the reduction of $\text{Fe}(\text{OH})_3$ (R4; between 0 and -4.2 cm depth) in the sediment porewater. The peak of ΣNO_3 in the overlying water may be associated with the oxidation of NH_4^+ (R7, R16). The presence of Fe concentrations close to the detection limit in the overlying water indicate its precipitation as $\text{Fe}(\text{OH})_3$ at the SWI (R10, R14), which is consistent with the significant enrichment in Fe_{ASC} in the surface sediment at this site (Table 2; Fig. 3a). The second diagenetic sequence occurs between -5.1 and -26.2 cm and involves the reduction of MnO_2 (from -5.1 to -7.6 cm), $\text{Fe}(\text{OH})_3$ (from -6.3 to -8.4 cm) and SO_4^{2-} (from -6.6 to -26.2 cm). The Mn produced by the oxidation of the OM (R3), Fe (R17) and $\Sigma\text{H}_2\text{S}$ (R18) diffuses upward where it is reoxidised by O_2 (R9) and/or ΣNO_3 (R13) at the SWI (consistent with the accumulation of MnO_2 in the surface sediment, Table 2), and downward, where it is co-precipitated with carbonates (R31). The Fe produced by the reduction of $\text{Fe}(\text{OH})_3$ by OM, NH_4^+ and/or $\Sigma\text{H}_2\text{S}$ (R4, R19, R20) diffuses upward where it is oxidised as $\text{Fe}(\text{OH})_3$ by MnO_2 (R17), and downward where it precipitates as Fe sulphides (R25). The presence of $\Sigma\text{H}_2\text{S}$ up to -3 cm depth can imply the precipitation of FeS (R25) in the same area where the Fe is produced. The presence of these two sequences can be explained by the activities of benthic organisms, which can mix the top few centimetres of sediment (sediment reworking) and/or modify the distribution of oxidized and reduced dissolved species near the SWI (bioirrigation). The occurrence of bioirrigation is illustrated by the O_2 peaks observed up to -0.4 cm under the SWI (Fig. 4).

4.2.3. Carbonate precipitation

The decreasing porewater salinity observed at the three sites (electronic annex EA-1) implies the decreasing concentrations of the major species of seawater. However, the decreases in the measured concentrations of SO_4^{2-} , Ca and Mg are greater than those expected from the decrease of salinity (calculated from the Dittmar's principle and Na concentrations, Fig. 9), suggesting a consumption of these elements in the sediments. As previously indicated, the consumption of SO_4^{2-} is associated to its reduction for OM mineralisation. The consumption of Ca and Mg is attributed to the precipitation of carbonate minerals reaching saturation ($SI \geq 0.5$ for calcite, aragonite, magnesite, dolomite and huntite). The ratios of the PROFILE-

calculated consumption rates of Mg to those of Ca are all greater than one ($R_{\text{net}} \text{Mg/Ca} = \sim 3$ for CB-O, ~ 5 for V-O and ~ 13 for CB-A), indicating the predominance of Mg precipitation relative to Ca and thus the preferential authigenic formation of Mg-rich carbonates. This carbonate precipitation may be responsible for the Mn consumption in the porewater by coprecipitation as previously proposed.

4.3. Reactions involved in the mobility of trace elements

4.3.1. Adsorption/Coprecipitation with $\text{Fe}(\text{OH})_3$

At site CB-A, the strong relationship between the depth intervals of $\text{Fe}(\text{OH})_3$ reduction and the remobilisation of As, Cr, Ni, Hg and MeHg (Fig. 7) is explained by the association of these species with $\text{Fe}(\text{OH})_3$ by adsorption and/or coprecipitation and by their remobilisation during the $\text{Fe}(\text{OH})_3$ reduction (R29). This is in accordance with the reported behaviour for As (Chaillou et al., 2003; Couture et al., 2010), Cr (Beck et al., 2008), Ni (Santos-Echeandia et al. 2009) and Hg/MeHg (Hollweg et al., 2009) in sediment.

The association of these trace elements by adsorption and/or coprecipitation with $\text{Fe}(\text{OH})_3$ (R29) is further supported at V-O by the As remobilisation that coincides with the depth interval where $\text{Fe}(\text{OH})_3$ is reduced (-4.8 and -9.5 cm) and in V-O and CB-O by the consumption of As, Ni and Cr below the SWI where precipitation of $\text{Fe}(\text{OH})_3$ occurs. The association with Fe cycle rather than Mn cycle is furthermore supported by the enrichments of As_{ASC} , Ni_{ASC} and, to a lesser extent, Cr_{ASC} in the surface sediments at the three sites as observed for Fe_{ASC} , while it is not the case for Mn_{ASC} that doesn't present any enrichment at site CB-O (Fig. 3). However, in the surface sediment at sites CB-A and CB-O, the extraction of a portion of As associated with the most reactive FeS or a portion of NiS (see below) can not be excluded.

Surprisingly, there is no remobilisation of As at the first peak of Fe at V-O. This observation also concerns ΣPO_4 that presents a similar geochemical behavior (Mucci et al., 2000), which is also only remobilised at the second peak of Fe (Fig. 5i). The major difference between these two Fe peaks is that the former occurs in the presence of high concentrations of $\text{Fe}(\text{OH})_3$ (Table 2, Fig. 3a). The absence of remobilisation of As and ΣPO_4 in this area implies that both species are associated with a less reactive phase of $\text{Fe}(\text{OH})_3$, which is reduced deeper in the sediment under more anoxic conditions, as proposed by Mucci et al. (2000). Another possibility is that under high concentrations of $\text{Fe}(\text{OH})_3$, the kinetics of the remobilisation of As and ΣPO_4 during the gradual dissolution of $\text{Fe}(\text{OH})_3$ are lower than their kinetics of readsorption onto the remaining $\text{Fe}(\text{OH})_3$ phase. A second surprising observation concerns the absence of Cr and Ni remobilisation during $\text{Fe}(\text{OH})_3$ reduction at site CB-O and of Cr and Ni at site V-O, which will be discussed in more detail in the following section.

The net reaction rates calculated for As remobilisation during $\text{Fe}(\text{OH})_3$ reduction ($4.4 \cdot 10^{-17} \text{ mol.cm}^{-3} \cdot \text{s}^{-1}$ and $0.6 \cdot 10^{-17} \text{ mol.cm}^{-3} \cdot \text{s}^{-1}$ for CB-A and V-O, respectively) or consumption during $\text{Fe}(\text{OH})_3$ precipitation ($-0.85 \cdot 10^{-17} \text{ mol.cm}^{-3} \cdot \text{s}^{-1}$ at V-O) are very similar to those reported using the same modelling approach in Canadian lakes: between $1.2 \cdot 10^{-17}$ and $7.1 \cdot 10^{-17} \text{ mol.cm}^{-3} \cdot \text{s}^{-1}$ for net remobilisation rates and between $-0.5 \cdot 10^{-17}$ and $-5.3 \cdot 10^{-17} \text{ mol.cm}^{-3} \cdot \text{s}^{-1}$ for net consumption rates (Couture et al. 2010).

4.3.2. Adsorption/Coprecipitation with MnO_2

Among the trace elements studied, only Co demonstrates a clear link with the Mn cycle, as previously reported in other aquatic environments (Canavan et al., 2007, Stockdale et al., 2010). Where $\Sigma\text{H}_2\text{S}$ is absent, areas of Co remobilisation correspond exactly to those where MnO_2 is reduced (Fig. 6; Fig. 8), supporting a strong coupling between Co and MnO_2 by adsorption and/or coprecipitation (R30). In the water column at site V-O, remobilised Co

diffuses downward and upward and is immediately trapped by readsorption and/or coprecipitation with MnO_2 . At CB-O, the strong Co remobilisation below the SWI could explain its depletion in the first centimetres of the solid phase sediment (Co_{ASC} , Fig. 3d). However, in presence of $\Sigma\text{H}_2\text{S}$, the affinity between Mn and Co appears to be somewhat altered.

4.3.3. Reactions with sulphides

At the three studied sites, excluding the particular case of Cr discussed in the last section, the trace elements diffuse downward from the depth interval where they are remobilised and are consumed in the deeper layers of sediment where sulphide is present (Fig. 6-8). In a sulphidic environment, the mobility of trace elements is greatly reduced in association with the formation of insoluble sulphide (e.g., Huerta-Diaz et al. 1998; Morse and Luther, 1999). The main processes involved in this immobilisation are direct metal sulphide precipitation or adsorption onto FeS/FeS_2 .

At the three sites, when $\Sigma\text{H}_2\text{S}$ is present, the thermodynamic calculations predict that the porewater concentrations of Ni and Co are at equilibrium with respect to $\text{NiS}\beta$ ($0 < \text{SI} < 1$) and $\text{CoS}\beta$ ($\text{SI} \sim 0$). Therefore, in the Berre Lagoon sediment, the precipitation of CoS and NiS clearly controls the mobility of Ni and Co in the sulphidic porewater (R27). The absence of Ni remobilisation during $\text{Fe}(\text{OH})_3$ reduction at CB-O and V-O indicates that the precipitation of NiS could kinetically dominate the reactions of remobilisation of Ni. In this case, the speciation of Ni in the solid phase would evolve during sediment burial from an adsorption onto $\text{Fe}(\text{OH})_3$ to an association with sulphides, without any significant remobilisation in the pore water. For Co, kinetic competition between its remobilisation during MnO_2 reduction (R30, cf section 4.3.2.) and CoS precipitation (R27) can be observed. Indeed, two areas are identified where MnO_2 reduction occurs in the presence of $\Sigma\text{H}_2\text{S}$: between 0 and -2.4 cm at CB-A (Fig. 7) and between -5 and -10 cm at V-O (Fig. 8). At CB-A, a high consumption rate of Co is observed at the same depth, indicating that the kinetics of precipitation of CoS clearly dominate those of Co remobilisation during MnO_2 reduction. At the site V-O, a different situation is observed because Co production rate is very low. In this latter case, we assumed that the remobilisation of Co dominates but is greatly limited by its reprecipitation in the form of CoS .

In contrast with Co and Ni, the PROFILE-predicted porewater consumptions in sulphidic sediments of As at sites CB-A and V-O and of Hg and MeHg at site CB-A cannot be attributed to the precipitation As or Hg sulphides nor to discrete mineral forms of As or Hg, since their porewater concentrations are all largely undersaturated with respect to the mineral tested (e.g., SI of realgar, orpiment and $\text{Mg}_3(\text{AsO}_4)_2 < -5$; SI of metacinnabar or cinnabar < -2). Alternatively, these patterns could be explained by the adsorption and/or coprecipitation of As and Hg/MeHg with FeS or FeS_2 (R28). According to the high proportion of MeHg ($73 \pm 22\%$) within the dissolved Hg_{tot} , we propose that the mobility of Hg in the sediment is primarily controlled by the adsorption of MeHg onto FeS . This could be further supported by the strong enrichment of MeHg in the solid phase at the SWI (Fig. 2k). The net rates obtained here for the adsorption and/or coprecipitation of As onto FeS/FeS_2 (between $-0.14 \cdot 10^{-17}$ and $-1.3 \cdot 10^{-17} \text{ mol.cm}^{-3} \cdot \text{s}^{-1}$) and of MeHg onto FeS ($-2.4 \cdot 10^{-21} \text{ mol.cm}^{-3} \cdot \text{s}^{-1}$) correspond well with those obtained, following the same method, by Couture et al. (2010) for As in Canadian lakes (between $-0.1 \cdot 10^{-17}$ and $-2.3 \cdot 10^{-17} \text{ mol.cm}^{-3} \cdot \text{s}^{-1}$) and by Merrit and Amirbahman (2007) for Hg_{tot} in the estuary of the Penobscot River (between $-7.5 \cdot 10^{-21}$ and $-14 \cdot 10^{-21} \text{ mol.cm}^{-3} \cdot \text{s}^{-1}$).

4.3.4. Cr Complexation with Dissolved Organic Matter

The behaviour of Cr at site V-O presents a particular pattern. Its high production in the porewater between -7 and -10 cm (Fig. 8) cannot be explained by one of the main diagenetic reactions because this is not a depth interval where a potential carrier phase (such as

Fe(OH)₃ or MnO₂) is dissolved. In contrast, the Cr peak in the porewater corresponds exactly to the depth of the peaks of Cr_{ASC} (Fig. 3e) or total Cr (Fig. 2i) in solid phase, indicating strong control of the porewater Cr concentrations by the solid/liquid phase equilibrium. Finally, the thermodynamic calculations predict a strong association of Cr with dissolved OM, with the presence of Cr complexes with humic acids (~ 40%) or fulvic acids (~ 20%). As previously shown, under anoxic conditions, Cr forms strong complexes with dissolved organic matter that greatly increase its mobility in porewater (e.g. Douglas et al. 1986; Beck et al., 2008). We therefore propose that the mobility of dissolved Cr in the absence of Fe(OH)₃ and MnO₂ in the sediments is dependent on its complexation with the dissolved OM in the porewater and the amount of solid phase Cr available.

The Cr profiles in the porewater at sites CB-A and CB-O cannot be associated with any of the previously identified main diagenetic reactions. However, in contrast with site V-O, there is no relation between the dissolved Cr concentrations in porewater and in particulate phase, most likely because of the lower Cr concentrations in the solid phase (Cr_{ASC} concentrations 2 and 8 times lower than at V-O; Fig. 3e) and lower DOC concentrations in the porewater (DOC concentrations 3 times lower than at V-O, Fig. 5I). At sites CB-A and CB-O, the mobility of Cr in the porewater is therefore expected to be very limited.

4.4. Fluxes at the SWI

The directions of the fluxes at the SWI are obtained from the concentration gradients of the dissolved species at the SWI (Table 5). When significant diffusive fluxes could be calculated by PROFILE, they are also reported.

4.4.1. Nutrients

Under oxic conditions at V-O, the high ΣNO_3 production a few centimetres above the SWI (Fig. 8) due to the nitrification reaction induces a flux to the water column. However, under slightly oxic conditions (CB-O), a ΣNO_3 flux to the sediment is predicted in relation to its consumption in the surface sediment (Table 5). These observations are consistent with the mechanisms of ΣNO_3 consumption/production that generally result in ΣNO_3 production from the sediment under oxic conditions and consumption in lower oxygen conditions (Deflandre et al., 2002; Belias et al., 2007, Middelburg and Levin, 2009; Emili et al., 2010). At CB-A, the absence of ΣNO_3 reflects complete consumption of this species, as expected given the anoxic character of the site.

Under anoxic conditions (CB-A), the sediment constitutes a source of NH_4^+ , ΣPO_4 as well as $\Sigma\text{H}_2\text{S}$ to the water column (Table 5), as reported in several different hypoxic/anoxic coastal systems (Belias et al. 2007; Emili et al., 2010; Mort et al., 2010). Under oxic conditions (CB-O and V-O), the $\Sigma\text{H}_2\text{S}$ flux disappeared due to its complete oxidation in the first cm of sediment (by Fe(OH)₃, MnO₂, ΣNO_3 and/or O₂), but positive ΣPO_4 and NH_4^+ fluxes are also observed at V-O in connection with the intense production of these species in the sediments (Fig. 8). However, their fluxes are much more lower than at site CB-A because of the oxidation of NH_4^+ (R7, R16, R19) and the accumulation of ΣPO_4 in the surface sediments by adsorption onto, or coprecipitation, with Fe(OH)₃ (R29) (Belias et al. 2007; Mort et al., 2010). At CB-O, these mechanisms eliminate the fluxes of ΣPO_4 and NH_4^+ to the water column (Table 5).

4.4.2. Fe, Mn and trace elements

The diffusive fluxes of trace elements at the SWI are observed to be highly dependent on the location of the diagenetic reactions inducing their remobilisation in the sediment. These reactions, previously identified are Fe(OH)₃ dissolution for As, Cr, Ni and Hg/MeHg and MnO₂ dissolution for Co.

The $\text{Fe}(\text{OH})_3$ reduction at the SWI at CB-A generates a positive flux of Fe and associated elements (As, Cr, Ni, Hg/MeHg, Table 5). However, only the remobilisation of As could be partially modelled by PROFILE (Fig. 7), which gives a minimum estimate of the flux as $> 3.5 \cdot 10^{-17} \text{ mol} \cdot \text{cm}^{-2} \cdot \text{s}^{-1}$. This is several orders of magnitude lower than the As flux calculated by Chaillou et al. (2003) in the Bay of Biscay, likely due to the strong trapping of As by adsorption/coprecipitation with FeS/FeS_2 and/or $\text{Fe}(\text{OH})_3$ in the surface sediment of Berre Lagoon.

At CB-O and V-O (Fig. 6, Fig. 8), the presence of O_2 in the water column and surface sediment leads to a deeper location of the reduction of $\text{Fe}(\text{OH})_3$ in the sediments (between -9 and -13 cm and between 0 and -4 cm, respectively) compared with CB-A. The remobilised Fe diffusing upwards is thus mostly or completely reoxidized as $\text{Fe}(\text{OH})_3$ in the surface sediments, which severely limits or prevents its diffusion to the water column (Table 5). The high concentration of reactive $\text{Fe}(\text{OH})_3$ in surface sediments, for which As, Cr and Ni have strong affinities, leads to their trapping in the surface sediments. For As and Cr, which are at low concentrations in both the water column and the porewater ($< 15 \text{ nM}$), no diffusive flux is identified. However for Ni, which has concentrations in the water column significantly higher than in the porewater, a negative flux is induced (Table 5). This observation is in contrast with common observations in coastal areas (Westerlund et al. 1986; Ciceri et al. 1992; Cheevaporn et al. 1995; Warnken et al. 2001; Santos-Echeandia et al., 2009), where exports of Ni from the sediment to the water column are generally reported. This is likely due to the high immobilisation efficiency of Ni by $\text{Fe}(\text{OH})_3$ and as NiS into the Berre Lagoon sediments.

The results for Cr at site V-O are peculiar. Despite the strong presence of $\text{Fe}(\text{OH})_3$ in the surface sediments, the high levels of dissolved organic ligands most likely explain the observed Cr release into the water column (Table 5) by promoting the mobility of Cr in the porewater and limiting its adsorption onto $\text{Fe}(\text{OH})_3$. This flux is in the range of those reported for other aquatic environments (Cheevaporn et al., 1995, Blasco et al., 2000). The primary source of Cr in the sediment is the highly enriched particulate phase at -10 cm (Fig. 2i), which corresponds to the Cr released into the lagoon in the 1960-1970s (Rigaud et al., 2011). Although over time this stock will gradually be buried, under similar conditions, the flux of Cr to the water column would most likely continue for several decades.

The MnO_2 reduction in surface sediments induces an Mn flux to the water column at the three sites (Table 5). However, these fluxes are very different, as the flux at CB-O is almost 100 times greater than at CB-A (not quantifiable at V-O). This difference is directly associated with the significantly higher MnO_2 reduction rate at CB-O in comparison with CB-A (approximately 70 times). At CB-O, 70% of the produced Mn is exported to the water column, versus 46% at CB-A.

At CB-O and V-O, positive fluxes of Co are observed in relation to its remobilisation during the MnO_2 dissolution (Table 5). The flux quantified at CB-O is in the range of those reported for Co in other coastal environments (Ciceri et al. 1992; Point et al. 2007; Santos-Echeandia et al., 2009) and results in the export of 60% of the Co remobilised under the SWI at CB-O. Under anoxic conditions (CB-A), despite the MnO_2 reduction at the SWI, the flux of Co occurs from the water column to the sediments (Table 5) due to the precipitation of CoS at the SWI.

The comparison of the calculated diffusive fluxes for those elements diffusing to the water column (Cr at V-O; As at CB-A and Co at CB-O) with the accumulation rate of the trace elements deposited on the sediments (estimated from the trace element concentrations in the surface sediments, the sedimentation rate and the wet density of the surface sediments), indicates highly efficient trapping of Cr at V-O sediments ($< 1\%$ of the deposited amount diffusing to the water column), but a significant amount (between 10 and 30%) of As (CB-A) and Co (CB-O) can be released to the water column following deposition.

In summary, the occurrence and intensity of the fluxes of trace elements to the water column are directly enhanced when their remobilisation reactions occur close to the SWI, and this situation is favoured under hypoxic/anoxic conditions.

4.5. Evaluation of the influence of the water column reoxygenation in the central basin

The anoxic conditions in the bottom water of the Central Basin of the Berre lagoon (Fig. 1) were quasi-perennial before 2000, largely dominant until 2005 and occurred approximately 35% of the time in 2009 (EDF, pers. com.). Anoxia has declined gradually with the establishment of regulations on the freshwater release from the hydroelectric power plant. In the future, it is possible that new regulations will continue to limit or stop these releases, inducing a partial or complete destratification of the water column and a probable increase of benthic water oxygen contents.

Our work enables to identify the main processes controlling trace element mobility in Berre lagoon sediments. The strong influence of the oxygen concentrations in the water column on both the intensity and direction of the fluxes at the SWI have been discussed. If we now consider the differences observed at site CB during anoxic conditions (CB-A) and oxic conditions (CB-O, $O_2 = 38\%$ O_2 saturation), it is possible to predict the influence of the reoxygenation of the water column in the central basin on the fluxes of nutrients and trace elements.

The three main species controlling the mobility of trace elements in the sediment of the Central Basin are ΣH_2S , Fe and Mn. Their evolution as conditions change from anoxic to oxic is thus particularly important to predict the evolution of the fluxes of trace elements. Under oxic conditions, these species are primarily oxidized in surface sediments causing 1) the precipitation of $Fe(OH)_3$ and MnO_2 at the SWI and the burial of their reduction zone and 2) the deepening of the occurrence of ΣH_2S in the sediment and thus the depth where sulphides can be involved in the immobilisation of trace elements.

An O_2 saturation of 38% in the water column (CB-O) seems to be sufficient to eliminate the upward flux of dissolved Fe from the sediment and to promote complete $Fe(OH)_3$ precipitation in the surface sediments. Therefore, we propose that reoxygenation of the water column $\geq 40\%$ of O_2 saturation will induce a reduction or an elimination of the upward fluxes of all of the species associated with the Fe cycle (As, Cr, Ni, Hg/MeHg and ΣPO_4) by their readsorption onto or, co-precipitation with, $Fe(OH)_3$ in the surface sediments. However, for Hg and MeHg, which were not measured under oxic conditions, the changes in their fluxes at the SWI under well oxygenated overlying water are more hypothetical and will largely depend on the competition between the kinetics of adsorption onto $Fe(OH)_3$ and the kinetics of remobilisation and/or methylation in the surface sediments.

Results for Mn are different with a calculated flux calculated higher in an oxygenated water column (38% O_2 saturation) than under anoxic conditions (Table 5). This difference may be attributed to the contrasted diagenetic behaviour of these two redox species: the MnO_2 reduction occurs for less reduced conditions than for $Fe(OH)_3$, and the oxidation kinetic of Fe is faster than that of Mn, which also requires microbial mediation (Burdige, 1993). The consequence is that an incomplete oxygenation ($\leq 40\%$ O_2 saturation) in the central basin will result in an intense MnO_2 reduction in the surface sediments and therefore a potential higher release of Co to the water column, in comparison with anoxic conditions. However, we can suppose that complete oxidation of Mn in the surface sediments may occur under complete oxygenation and may eliminate the flux of Mn and Co to the water column.

Water column oxygenation of 38% O_2 saturation results in the oxidation of NH_4^+ and the decrease or elimination of its flux from the sediments at site CB (Table 5). However, it also results in the formation of ΣNO_3 in the water column under high oxygen concentrations, as observed at V-O (Fig. 8), which can contribute to the eutrophication of the system.

All of these predictions are based only on the biogeochemical processes studied in this work. This does not consider the influence of sediments recolonisation by macrobenthic organisms that is expected upon complete reoxygenation of the bottom water. In such case, the bioturbation processes would significantly disturb the locations of the main diagenetic reactions in the surface sediments and thus the fluxes of the trace elements and nutrients, as evidenced by the relative complexity of their profiles at V-O (Morse and Eldridge, 2007; Middelburg and Levin, 2009). Consequently, predictions previously made are thus valid in the case of a likely scenario of decrease but persistence of anoxic events in the Central Basin that will prevent sediment recolonization by macrofauna.

5. Conclusion

Coupling multiple methodological approaches (microelectrodes, peepers, sediment cores, chemical extractions, thermodynamic calculations and steady-state transport-reaction modelling) in a Mediterranean lagoon exhibiting both contaminated sediments and contrasting oxygen concentrations in the overlying water allowed for the identification of processes that control the mobility and fluxes of trace elements (As, Co, Cr, Ni, Hg, MeHg) and nutrients (ΣNO_3 , NH_4^+ , ΣPO_4) in the sediment under different oxygen contents in the water column.

The main diagenetic reactions identified in the control of the mobility and fluxes of trace elements in sediments are the adsorption onto, or co-precipitation with, Fe oxy-hydroxides (As, Cr, Ni, Hg/MeHg), Mn oxy-hydroxydes (Co) and FeS/FeS₂ (As, MeHg) or the direct precipitation as metal sulphides (Co, Ni). The remobilisation of trace elements occurs principally when Fe and Mn oxy-hydroxydes are reduced, but such remobilisation can be partially or totally prevented in the presence of sulphide depending on the kinetic competitions. The fluxes of trace elements to the overlying water are then essentially dependent on the location of these remobilisation reactions in the sediment column, which were shown to be primarily dependent on the oxygenation of the water column.

Under anoxic conditions in the overlying water, reduction of Fe and Mn oxy-hydroxydes and SO_4^{2-} (and thus $\Sigma\text{H}_2\text{S}$ production) occurs directly at the SWI. This leads to a kinetic competition between remobilisation and trapping reactions at the SWI. This resulted in the diffusion of Fe, Mn, Cr, Ni, As and Hg/MeHg as well as $\Sigma\text{H}_2\text{S}$, NH_4^+ and ΣPO_4 to the water column. However for Co, the precipitation as CoS at the SWI dominates the Co remobilisation during Mn oxy-hydroxydes reduction, preventing its diffusion to the overlying water. At the same site under oxic conditions, the deepening of depth interval where Fe oxy-hydroxides are reduced in the sediment induces a decrease or elimination of the upward fluxes of associated trace elements to the water column. However, the disappearance of $\Sigma\text{H}_2\text{S}$ in surface sediment prevents the immobilisation of Co (associated with Mn and not Fe) and leads, in this case, to an increase in its flux to the overlying water column. At the site exhibiting perennial oxic conditions in water column, the sediments are a source of ΣNO_3 , Fe, Mn, Co and Cr and to a lesser extent of ΣPO_4 and NH_4^+ , but still are a sink for Ni. The behaviour of Cr in the sulphidic sediment is peculiar because its remobilisation has been observed to be decoupled from the main diagenetic reactions and primarily controlled by its complexation with dissolved organic matter in the porewater.

A potential future reoxygenation of the water column in the central basin of the Berre Lagoon will likely decrease or eliminate the release of trace elements and $\Sigma\text{H}_2\text{S}$, NH_4^+ and ΣPO_4 to the water column. However, if the oxygenation of the water column is not complete, the fluxes of Mn and Co will likely be much higher than those prevailing under the current anoxic

conditions. Finally, the improvement of oxygenation of the water column will likely increase the fluxes of ΣNO_3 .

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Tables

Table 1: Location and characteristic of sampling sites. The physicochemical parameters reported correspond to the deepest (about 10-20 cm above the SWI) measure carried out for each site in the water column.

	CB-A	CB-O	V-O
Location	05°05.845 E / 43°26.529 N		5°12.166 E / 43°28.032 N
Water column depth (m)		9.0	4.5
Sedimentation rate (cm.a ⁻¹) ^a		0.28	0.34
Temperature (°C)	21	18	21
Salinity	35.5	32.6	20.8
pH	7.45	7.65	7.35
O ₂ (mg.L ⁻¹) / (% saturation)	<0.2 / <3	2.9 / 38	4.6 / 60
Eh (mV/NHE)	40	330	380

(a) from Salençon, Pers. Com..

Table 2: Average concentrations of Fe_{ASC} and Mn_{ASC} reactive minerals (as µg of Fe or Mn per g of sediment or % of total Fe or Mn) in sediments over the depth interval identified using ascorbate extractions.

Reactive minerals	Depth interval (cm)	CB-A		CB-O			V-O		
		C (µg.g ⁻¹)	% _{tot}	Depth interval (cm)	C (µg.g ⁻¹)	% _{tot}	Depth interval (cm)	C (µg.g ⁻¹)	% _{tot}
Fe _{ASC}	0-3	1960	9.8	0-3	2110	11.3	0-3	5230	23.9
Mn _{ASC}	0-2	19.4	5.5	0-2	9.5	2.7	0-3	83.3	21.9

Table 3: Main diagenetic reactions and reactions involving trace elements. Me^{2+} , $\text{Fe}(\text{OH})_3$ and MnO_2 represent the trace elements, the Fe and Mn oxy-hydroxides respectively. Me- $\text{Fe}(\text{OH})_3$, Me- MnO_2 , Me- FeS , Me FeS_2 and Me-OM represent trace elements association on particulate phases by adsorption or coprecipitation without specifying the nature. 1) Wang and Van Cappellen, 1996 ; 2) Canavan et al., 2006 ; 3) Anschutz et al., 2000 ; 4) Hulth, 1999 ; 5) Boudreau, 1997.

No.	Description	Reactions	References
<i>Primary redox reactions</i>			
R1	Aerobic respiration	$\text{OM} + x\text{O}_2 + (-y + 2z)\text{HCO}_3^- \rightarrow (x - y + 2z)\text{CO}_2 + y\text{NH}_4^+ + z\text{HPO}_4^{2-} + (x + 2y + 2z)\text{H}_2\text{O}$	1, 2
R2	Denitrification	$\text{OM} + 0.8x\text{NO}_3^- \rightarrow 0.4x\text{N}_2 + (0.2x - y + 2z)\text{CO}_2 + (0.8x + y - 2z)\text{HCO}_3^- + y\text{NH}_4^+ + z\text{HPO}_4^{2-} + (0.6x - y + 2z)\text{H}_2\text{O}$	1, 2
R3	MnO_2 reduction	$\text{OM} + 2x\text{MnO}_2 + (3x + y - 2z)\text{CO}_2 + (x + y - 2z)\text{H}_2\text{O} \rightarrow 2x\text{Mn}^{2+} + (4x + y - 2z)\text{HCO}_3^- + y\text{NH}_4^+ + z\text{HPO}_4^{2-}$	1, 2
R4	$\text{Fe}(\text{OH})_3$ reduction	$\text{OM} + 4x\text{Fe}(\text{OH})_3 + (7x + y - 2z)\text{CO}_2 \rightarrow 4x\text{Fe}^{2+} + (8x + y - 2z)\text{HCO}_3^- + y\text{NH}_4^+ + z\text{HPO}_4^{2-} + (3x + y - 2z)\text{H}_2\text{O}$	1, 2
R5	Sulphates reduction	$\text{OM} + 0.5x\text{SO}_4^{2-} + (y - 2z)\text{CO}_2 + (y - 2z)\text{H}_2\text{O} \rightarrow 0.5x\text{H}_2\text{S} + (x + y - 2z)\text{HCO}_3^- + y\text{NH}_4^+ + z\text{HPO}_4^{2-}$	1, 2
R6	Methanogenesis	$\text{OM} + (y - 2z)\text{H}_2\text{O} \rightarrow 0.5x\text{CH}_4 + (0.5x - y + 2z)\text{CO}_2 + (y - 2z)\text{HCO}_3^- + y\text{NH}_4^+ + z\text{HPO}_4^{2-}$	1, 2
<i>Secondary redox reactions</i>			
R7	NH_4^+ oxidation by O_2	$\text{NH}_4^+ + 2\text{O}_2 + 2\text{HCO}_3^- \rightarrow \text{NO}_3^- + 2\text{CO}_2 + 3\text{H}_2\text{O}$	1,2,3,4
R8	H_2S oxidation by O_2	$\text{H}_2\text{S} + 2\text{O}_2 + 2\text{HCO}_3^- \rightarrow \text{SO}_4^{2-} + 2\text{CO}_2 + 2\text{H}_2\text{O}$	1,2
R9	Mn^{2+} oxidation by O_2	$\text{Mn}^{2+} + 1/2\text{O}_2 + 2\text{HCO}_3^- \rightarrow \text{MnO}_2 + 2\text{CO}_2 + \text{H}_2\text{O}$	1,2,3
R10	Fe^{2+} oxidation by O_2	$\text{Fe}^{2+} + 1/4\text{O}_2 + 2\text{HCO}_3^- + 1/2\text{H}_2\text{O} \rightarrow \text{Fe}(\text{OH})_3 + 2\text{CO}_2$	1,2
R11	CH_4 oxidation par O_2	$\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$	1,2
R12	FeS oxidation by O_2	$\text{FeS} + 2\text{O}_2 \rightarrow \text{Fe}^{2+} + \text{SO}_4^{2-}$	1,2
R13	Mn^{2+} oxidation by NO_3^-	$5\text{Mn}^{2+} + 2\text{NO}_3^- + 4\text{H}_2\text{O} \rightarrow 5\text{MnO}_2 + \text{N}_2 + 8\text{H}^+$	4
R14	Fe^{2+} oxidation by NO_3^-	$5\text{Fe}^{2+} + \text{NO}_3^- + 12\text{H}_2\text{O} \rightarrow 5\text{Fe}(\text{OH})_3 + 1/2\text{N}_2 + 9\text{H}^+$	4
R15	FeS oxidation by NO_3^-	$5\text{FeS} + 8\text{NO}_3^- + 8\text{H}^+ \rightarrow 5\text{Fe}^{2+} + 5\text{SO}_4^{2-} + 4\text{N}_2 + 4\text{H}_2\text{O}$	4
R16	NH_4^+ oxidation by MnO_2	$\text{NH}_4^+ + 4\text{MnO}_2 + 6\text{H}^+ \rightarrow 4\text{Mn}^{2+} + \text{NO}_3^- + 5\text{H}_2\text{O}$ ou $3\text{MnO}_2 + 2\text{NH}_4 + 4\text{H}^+ = 3\text{Mn}^{2+} + \text{N}_2 + 6\text{H}_2\text{O}$	3,4
R17	Fe^{2+} oxidation by MnO_2	$2\text{Fe}^{2+} + \text{MnO}_2 + 2\text{HCO}_3^- + 2\text{H}_2\text{O} \rightarrow 2\text{Fe}(\text{OH})_3 + \text{Mn}^{2+} + 2\text{CO}_2$	1,2
R18	H_2S oxidation by MnO_2	$\text{H}_2\text{S} + 4\text{MnO}_2 + 6\text{CO}_2 + 2\text{H}_2\text{O} \rightarrow 4\text{Mn}^{2+} + 6\text{HCO}_3^- + \text{SO}_4^{2-}$	2
R19	NH_4^+ oxidation by $\text{Fe}(\text{OH})_3$	$\text{NH}_4^+ + 8\text{Fe}(\text{OH})_3 + 14\text{H}^+ \rightarrow 8\text{Fe}^{2+} + \text{NO}_3^- + 21\text{H}_2\text{O}$	3
R20	H_2S oxidation by $\text{Fe}(\text{OH})_3$	$\text{H}_2\text{S} + 8\text{Fe}(\text{OH})_3 + 14\text{CO}_2 \rightarrow 8\text{Fe}^{2+} + 14\text{HCO}_3^- + \text{SO}_4^{2-} + 6\text{H}_2\text{O}$	2
R21	CH_4 oxidation by SO_4^{2-}	$\text{CH}_4 + \text{CO}_2 + \text{SO}_4^{2-} \rightarrow 2\text{HCO}_3^- + \text{H}_2\text{S}$	1,2
<i>Precipitation reactions (non-redox)</i>			
R22	CaCO_3 precipitation	$\text{Ca}^{2+} + \text{HCO}_3^- \rightarrow \text{CaCO}_3 + \text{H}^+$	5
R23	MnCO_3 precipitation	$\text{Mn}^{2+} + 2\text{HCO}_3^- \rightarrow \text{MnCO}_3 + \text{CO}_2 + \text{H}_2\text{O}$	1
R24	FeCO_3 precipitation	$\text{Fe}^{2+} + 2\text{HCO}_3^- \rightarrow \text{FeCO}_3 + \text{CO}_2 + \text{H}_2\text{O}$	1
R25	FeS precipitation	$\text{Fe}^{2+} + 2\text{HCO}_3^- + \text{H}_2\text{S} \rightarrow \text{FeS} + 2\text{CO}_2 + 2\text{H}_2\text{O}$	1,2
R26	Pyrite precipitation	$\text{FeS} + \text{H}_2\text{S} \rightarrow \text{FeS}_2 + \text{H}_2$	2
<i>Reactions involving trace elements</i>			
R27	Trace elements sulphides precipitation	$\text{Me}^{2+} + \text{H}_2\text{S} \rightarrow \text{MeS} + 2\text{H}^+$	
R28	Adsorption/coprecipitation on FeS/FeS ₂	$\text{Me}^{2+} + \text{FeS} \rightarrow \text{Me-FeS}$ or $\text{Me}^{2+} + \text{FeS}_2 \rightarrow \text{Me-FeS}_2$	
R29	Adsorption/coprecipitation on $\text{Fe}(\text{OH})_3$	$\text{Me}^{2+} + \text{Fe}(\text{OH})_3 \rightarrow \text{Me-Fe}(\text{OH})_3$	
R30	Adsorption/coprecipitation on MnO_2	$\text{Me}^{2+} + \text{MnO}_2 \rightarrow \text{Me-MnO}_2$	
R31	Adsorption/coprecipitation on CO_3	$\text{Me}^{2+} + \text{CO}_3 \rightarrow \text{Me-CO}_3$	
R32	Adsorption on OM	$\text{Me}^{2+} + \text{OM} \rightarrow \text{Me-OM}$	

With $\text{OM} = (\text{CH}_2\text{O})_x(\text{NH}_3)_y(\text{H}_3\text{PO}_4)_z$ and x, y and z are the molar coefficients of C, N and P elements.

Table 4: Main diagenetic reactions and their depth intervals

Main diagenetic reactions	Depth intervals (cm)		
	CB-O	CB-A	V-O
O ₂ consumption	0/-0.045	no	0/-0.19
Denitrification	0/-5.7	no	+5.0/-2
MnO ₂ reduction	0/-2.4	0/-3.7	+1/+3 ; -5.1/-7.6
Fe(OH) ₃ reduction	-9.5/-11.9	+1/0	0/-4.2 ; -6.3/-8.4
SO ₄ ²⁻ reduction	-14.3/-28.6	0/-5.0	-6.6/-26.2
FeS/FeS ₂ precipitation	-11.9/-28.6	0/-2.0	(-3?)-5.9/-10.6
Carbonate precipitation	-10/-28.6	-2/-30	-10/-26.2

Table 5: Synthesis of fluxes at the SWI in sites CB-O, CB-A and V-O. The arrows represent the direction of the fluxes. When significant, the diffusive fluxes (mol.cm⁻².s⁻¹), quantified from the concentration gradients modeled below the SWI by PROFILE are reported. A negative value indicates a flux from the overlying water to the sediment, a positive value indicates a flux from the sediment to the overlying water.

	CB-O	CB-A	V-O
O ₂	↓(-8.7.10 ⁻¹¹)		↓(-3.2.10 ⁻¹¹)
NO ₃ ⁻	↓(-1.89.10 ⁻¹⁴)		↑
∑NH ₄ ⁺	~0	↑(6.9.10 ⁻¹³)	↑(6.2.10 ⁻¹³)
∑PO ₄	~0	↑(2.3.10 ⁻¹⁴)	↑(0.1.10 ⁻¹⁴)
∑H ₂ S	~0	↑(4.5.10 ⁻¹³)	0
Fe	~0	↑	↑(0.56.10 ⁻¹⁴)
Mn	↑(1.17.10 ⁻¹³)	↑(0.02.10 ⁻¹³)	↑
As	~0	↑(>3.5.10 ⁻¹⁷)	~0
Co	↑(8.20.10 ⁻¹⁷)	↓(-0.41.10 ⁻¹⁷)	↑
Cr	~0	↑	↑(1.1.10 ⁻¹⁷)
Ni	↓(-8.7.10 ⁻¹⁸)	↑	↓(-1.7.10 ⁻¹⁸)
Hg	n.d.	↑	n.d.
MeHg	n.d.	↑	n.d.

Figures

Figure 1: Study area and sampling sites. The area of the Central Basin subject to anoxic events and presenting no macrobenthic organism, is reported.

Figure 2: Total concentration profiles of (a-f) selected majors constituents (TOC, N_{tot} , TIC, Fe, Mn and S) and (g-l) trace elements (As, Co, Cr, Hg, MeHg and Ni) in sediment cores from sites CB-A, CB-O and V-O. When available, the natural geochemical background (NGB, mean \pm 1 standard deviation) for trace elements for the area obtained from the deepest part of long sediment cores (Rigaud et al., 2011) is reported.

Figure 3: Profile of Fe, Mn, As, Co, Cr and Ni concentrations extracted by ascorbate in CB-A, CB-O and V-O sediments cores.

Figure 4: Oxygen profiles obtained by ex-situ microelectrodes measurements with overlying water oxygenation adjusted to in-situ conditions. The horizontal dashed line represents the sediment/water interface.

Figure 5: Concentration profiles of major species (a-l) and trace elements (m-r) in pore and overlying waters from peepers on sites CB-A, CB-O and V-O obtained by peepers. The horizontal dashed line represents the sediment/water interface.

Figure 6: A) concentration profiles of ΣNO_3 , Mn, Fe, SO_4^{2-} , ΣH_2S , Co, Cr and Ni obtained by peeper in the site CB-O. The fit of experimental values, areas and net rates of reactions obtained using PROFILE are also reported. B) Schematic representation of areas and depth integrated rates of production/consumption reactions of dissolved species in porewaters obtained by PROFILE. The thickness of the forms is proportional to the concentration of the species along the profile. The reactions involved in these processes are reported by their numbers given in Table 3. Grey color indicates production, white is for consumption (see text for detail). The significant diffusive fluxes at the SWI are also reported using arrows. As profile is not reported here since PROFILE could not determine statistically significant zones where it could be either produced or consumed from the porewaters owing to the high level of noise in the profile (Fig. 5m).

Figure 7: A) concentration profiles of Mn, Fe, SO_4^{2-} , ΣH_2S , As, Co, Cr, Ni, Hg_{tot} and MeHg obtained by peeper in the site CB-A. The fit of experimental values, areas and net rates of reactions obtained using PROFILE are also reported. B) Schematic representation of areas and depth integrated rates of production/consumption reactions of dissolved species in porewaters obtained by PROFILE. The thickness of the forms is proportional to the concentration of the species along the profile. The reactions involved in these processes are reported by their numbers given in Table 3. Grey color indicates production, white is for consumption (see text for detail). The significant diffusive fluxes at the SWI are also reported using arrows. The profiles of Hg and MeHg correspond to the average of the three replicates presented in the Fig. 5q-r.

Figure 8: A) concentration profiles of ΣNO_3 , Mn, Fe, SO_4^{2-} , ΣH_2S , As, Co, Cr and Ni obtained by peeper in the site V-O. The fit of experimental values, areas and net rates of reactions obtained using PROFILE are also reported. B) Schematic representation of areas and depth integrated rates of production/consumption reactions of dissolved species in porewaters obtained by PROFILE. The thickness of the forms is proportional to the concentration of the species along

the profile. The reactions involved in these processes are reported by their numbers given in Table 3. Grey color indicates production, white is for consumption (see text for detail). The significant diffusive fluxes at the SWI are also reported using arrows.

Figure 9: Concentration profiles of Ca, Mg and SO_4^{2-} measured and calculated using the Dittmar's principle and Na concentration in sites CB-O, CB-A and V-O. The horizontal dotted line represents the SWI. The abscissa scale is adjusted for better viewing.

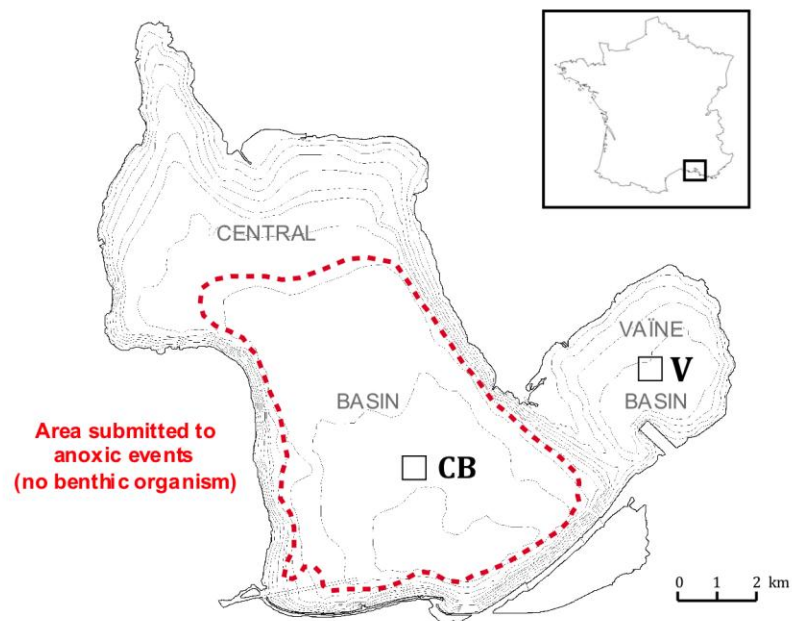


Fig. 1

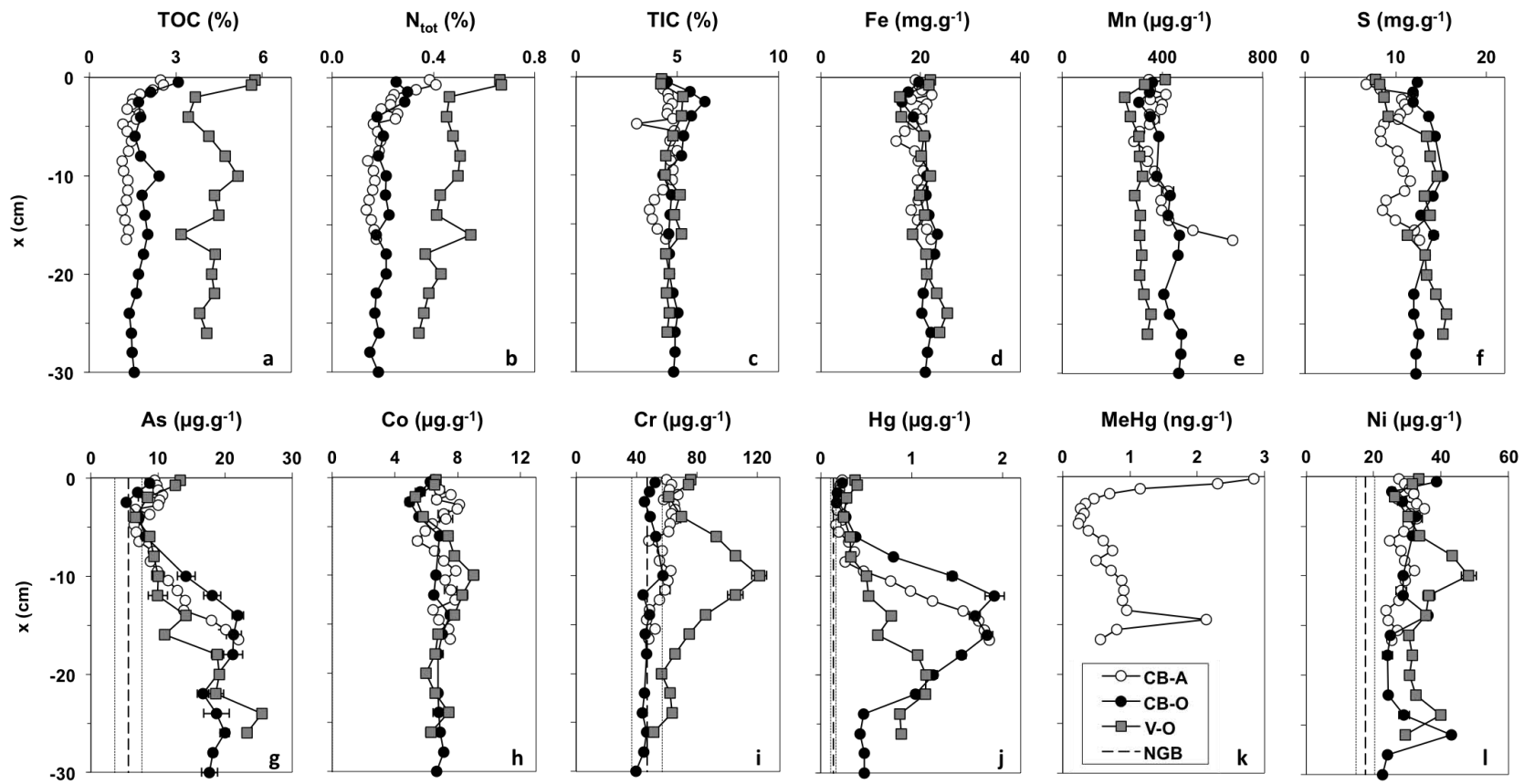


Fig. 2

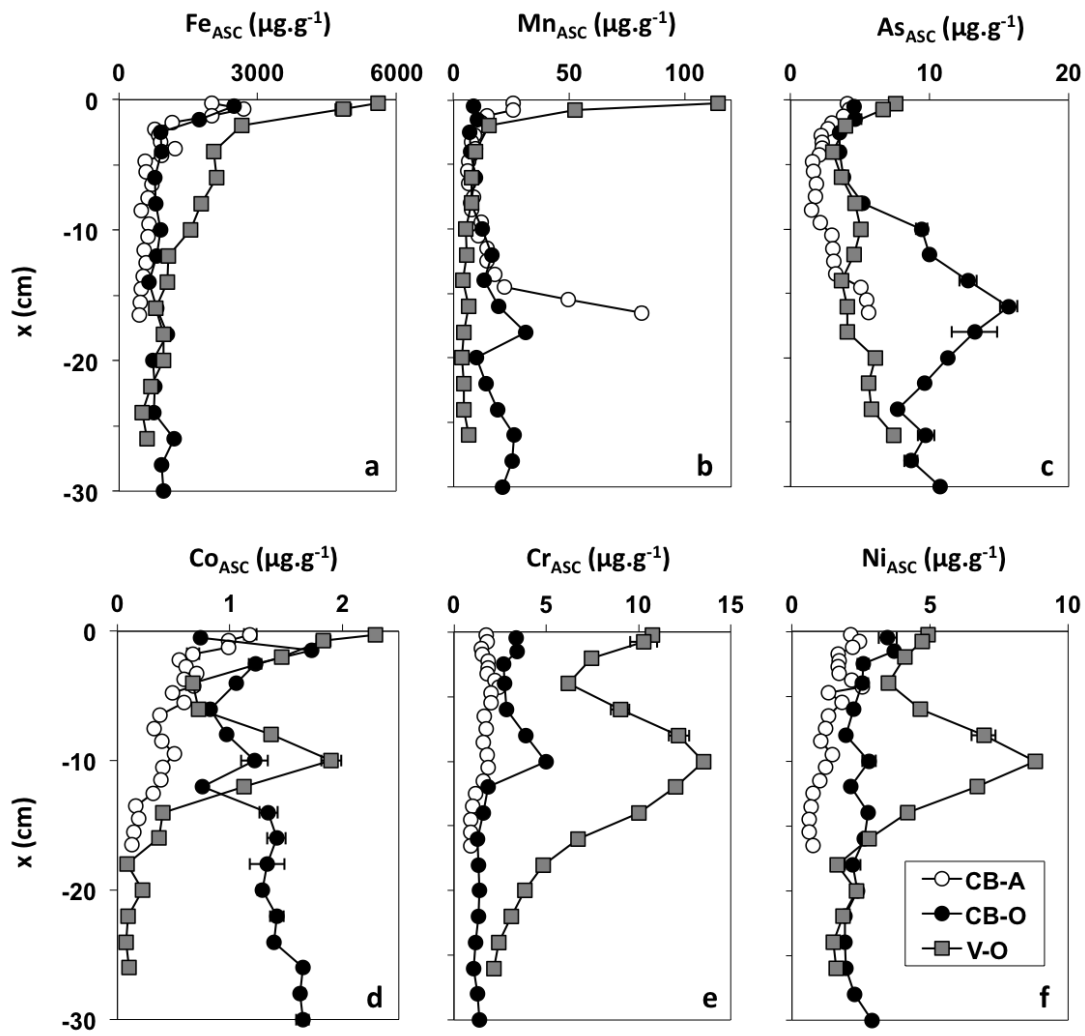


Fig. 3

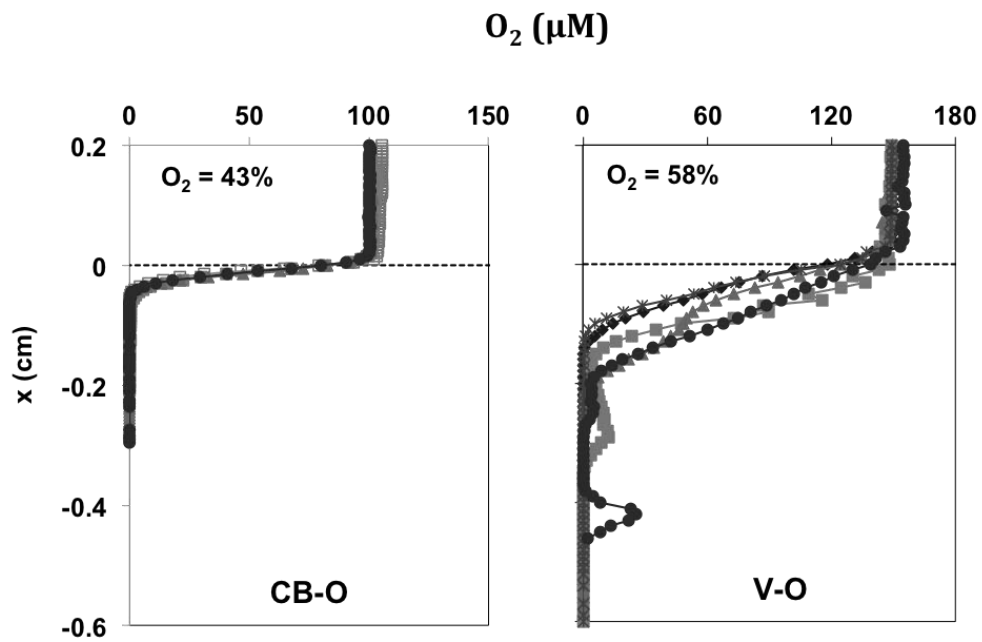


Fig. 4

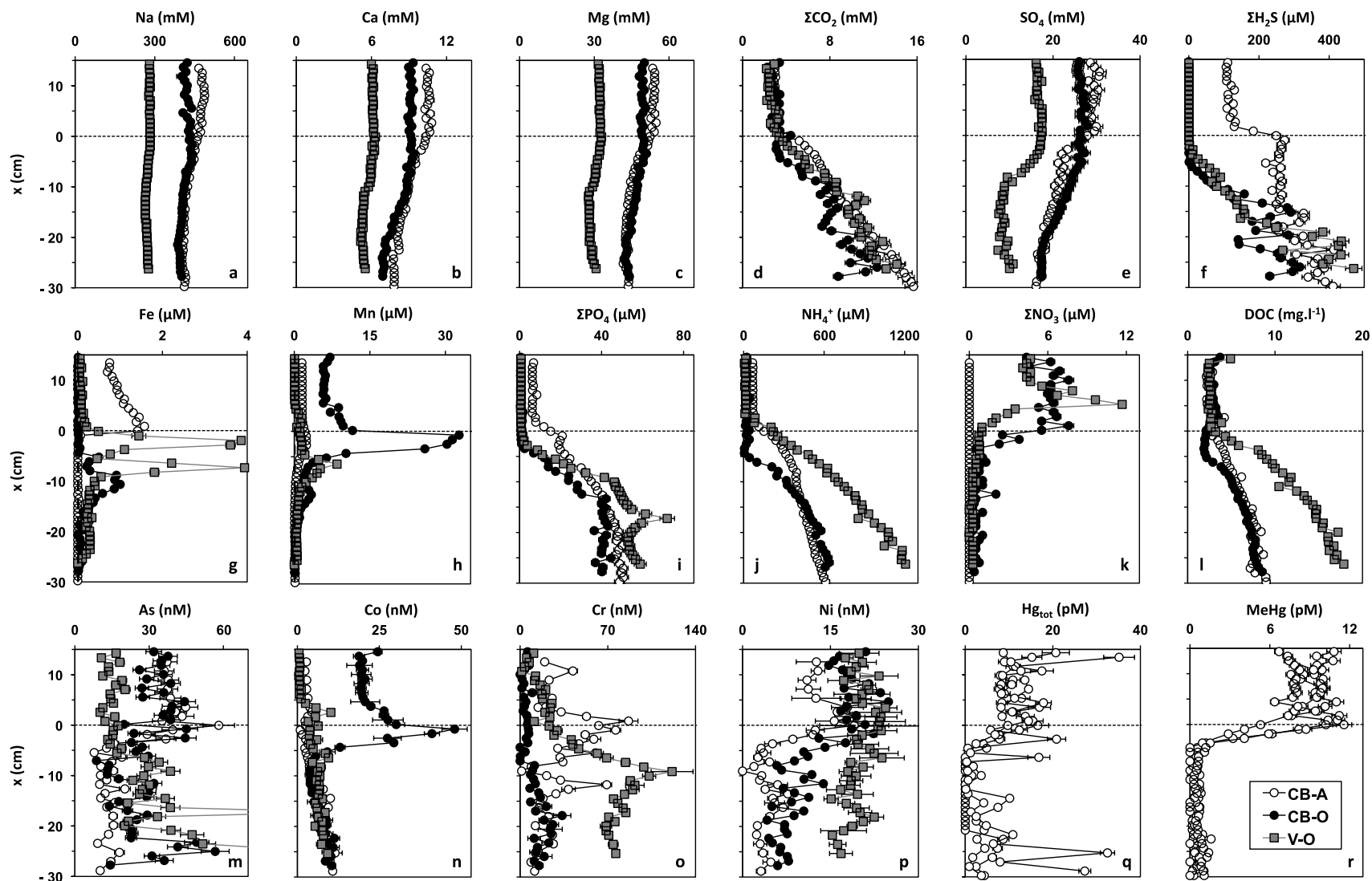


Fig. 5

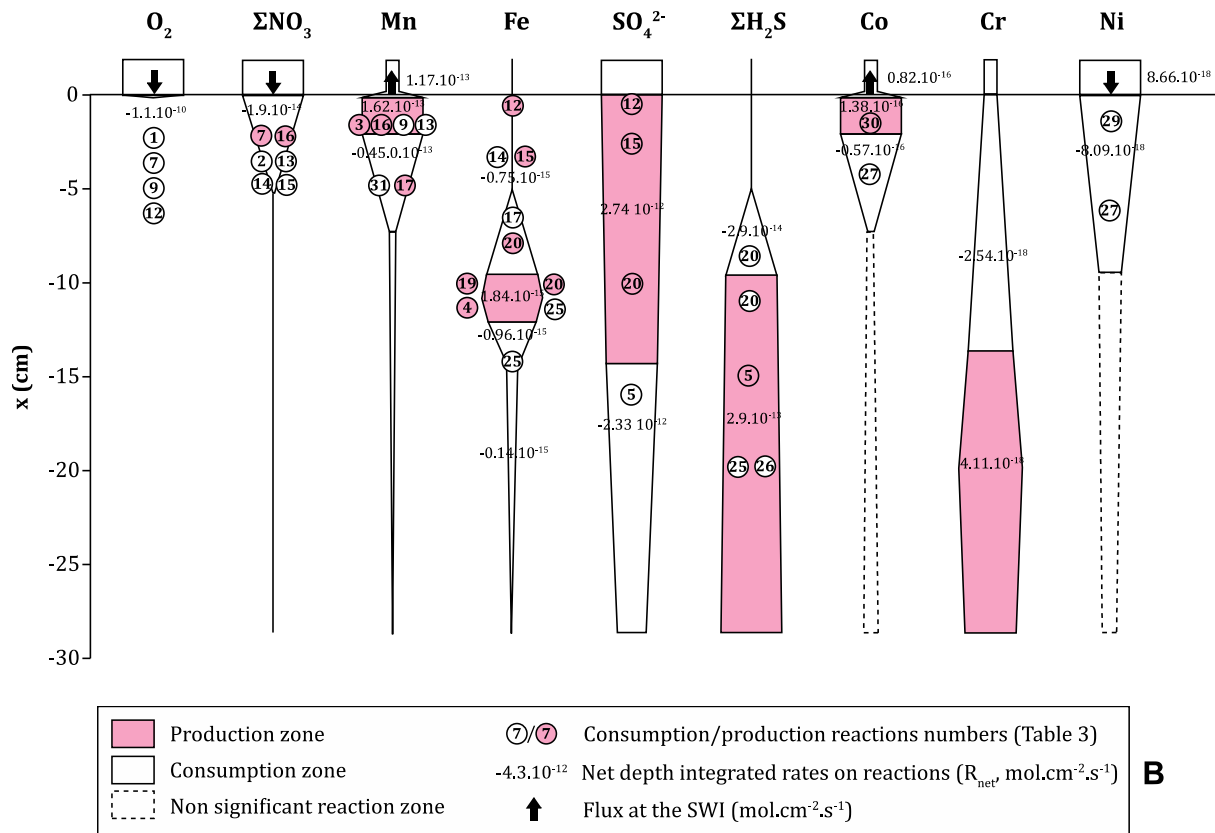
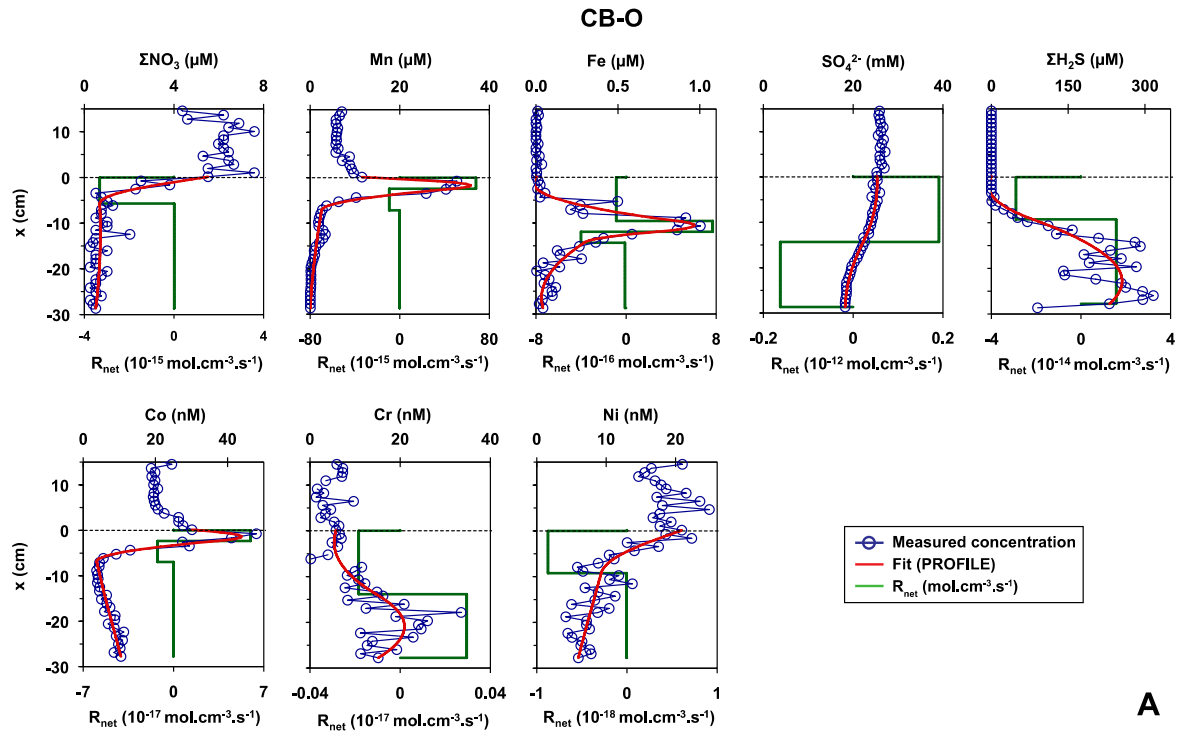


Fig. 6

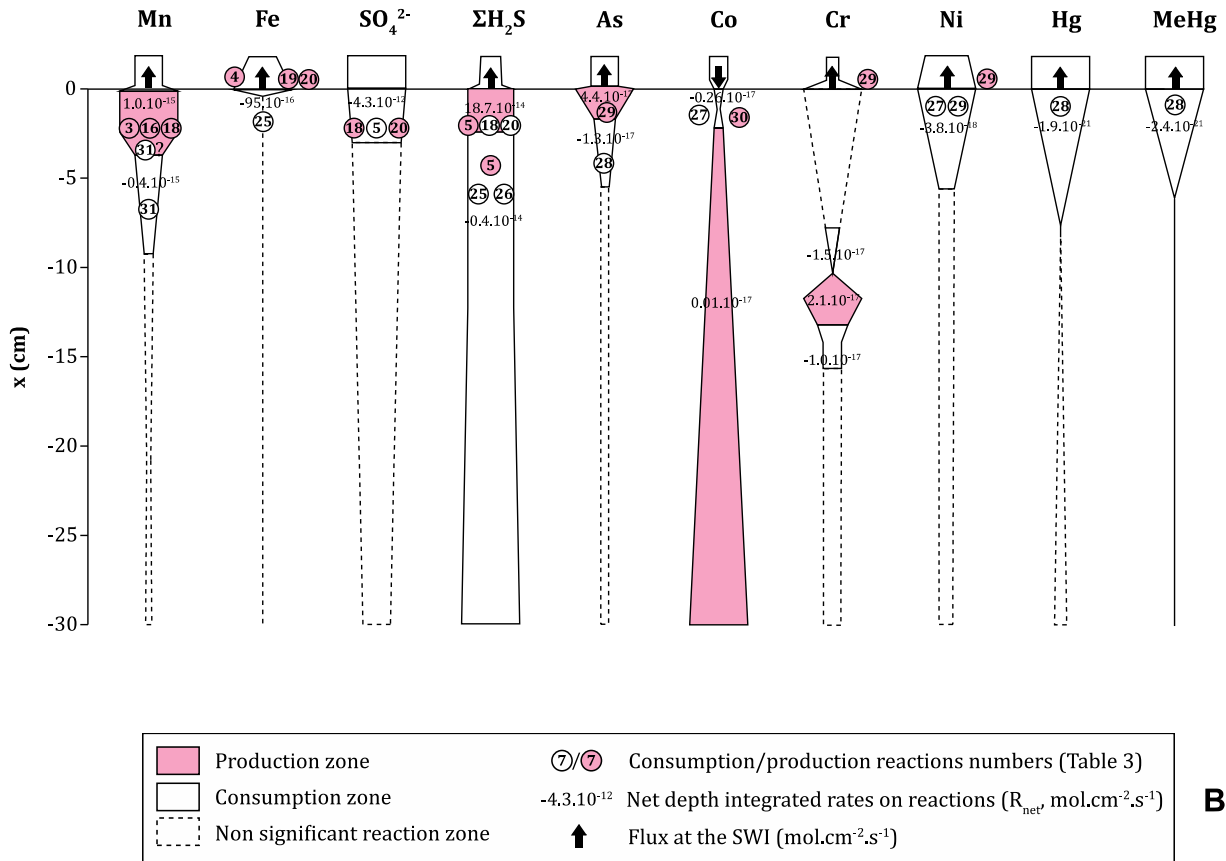
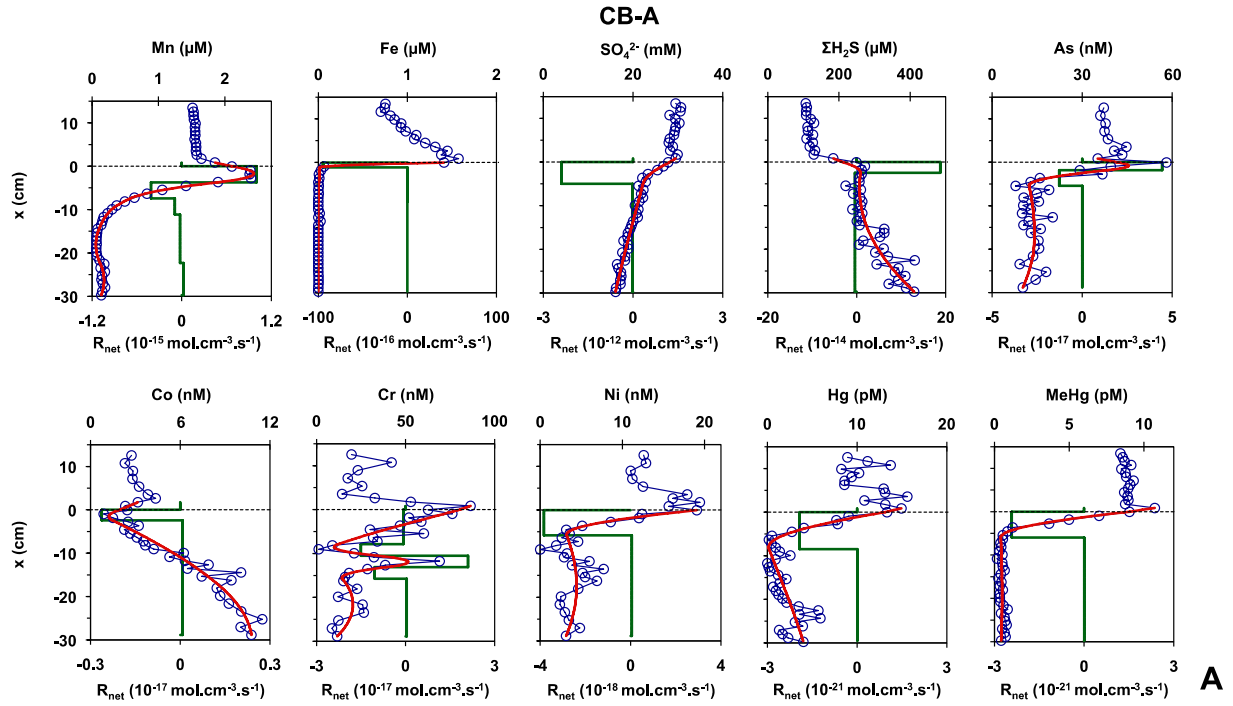


Fig. 7

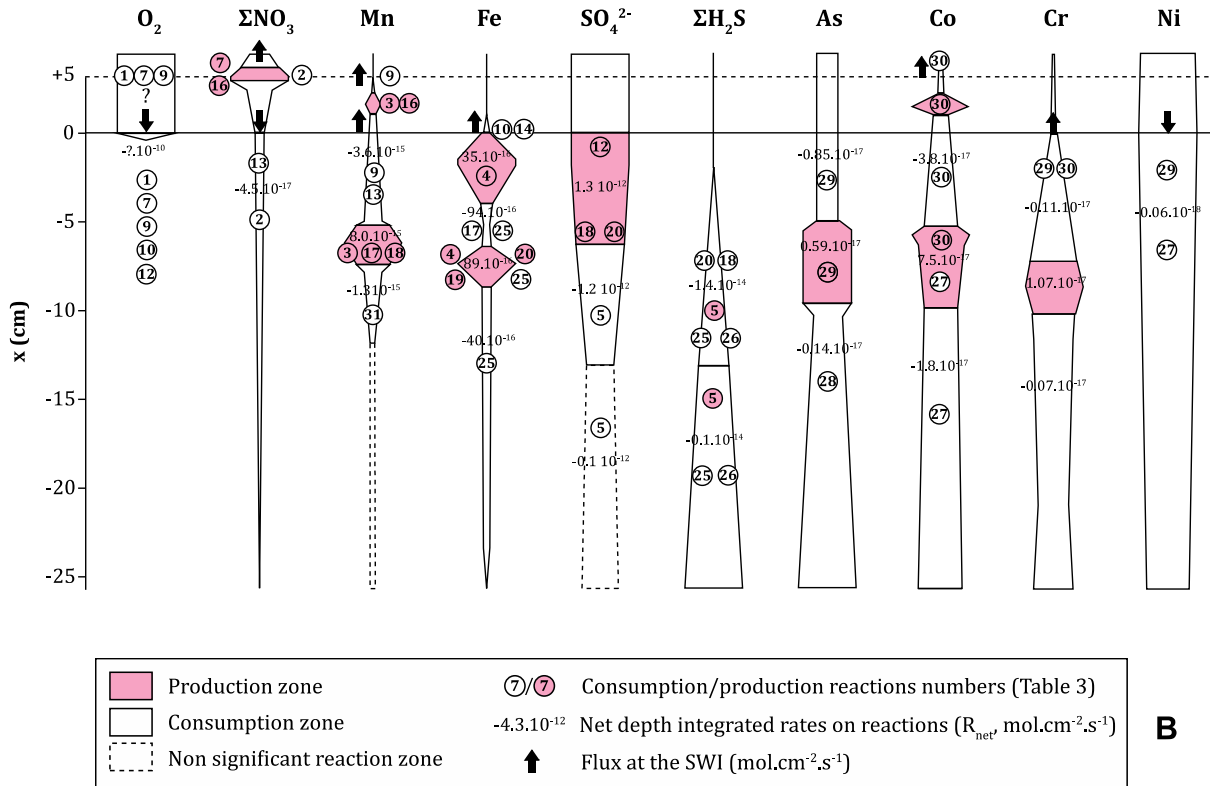
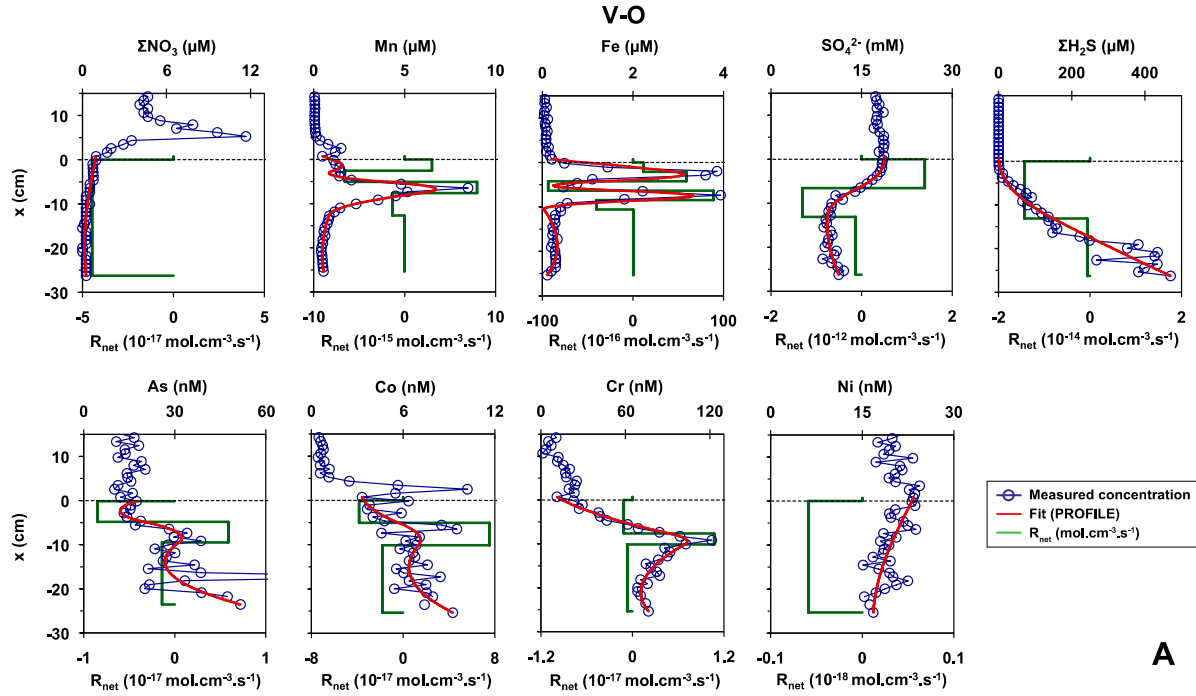


Fig. 8

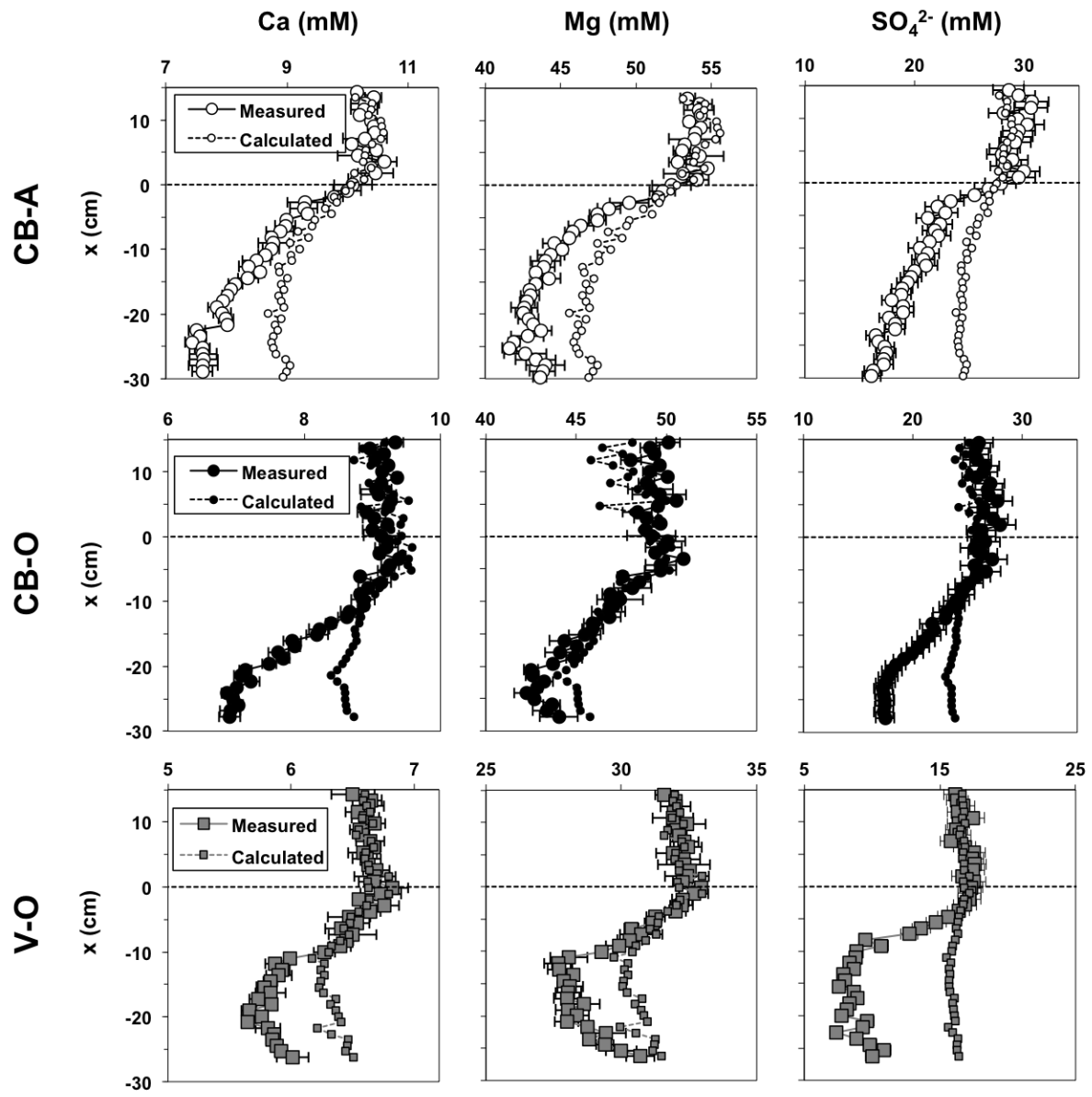


Fig. 9

EA-1: Supplementary data

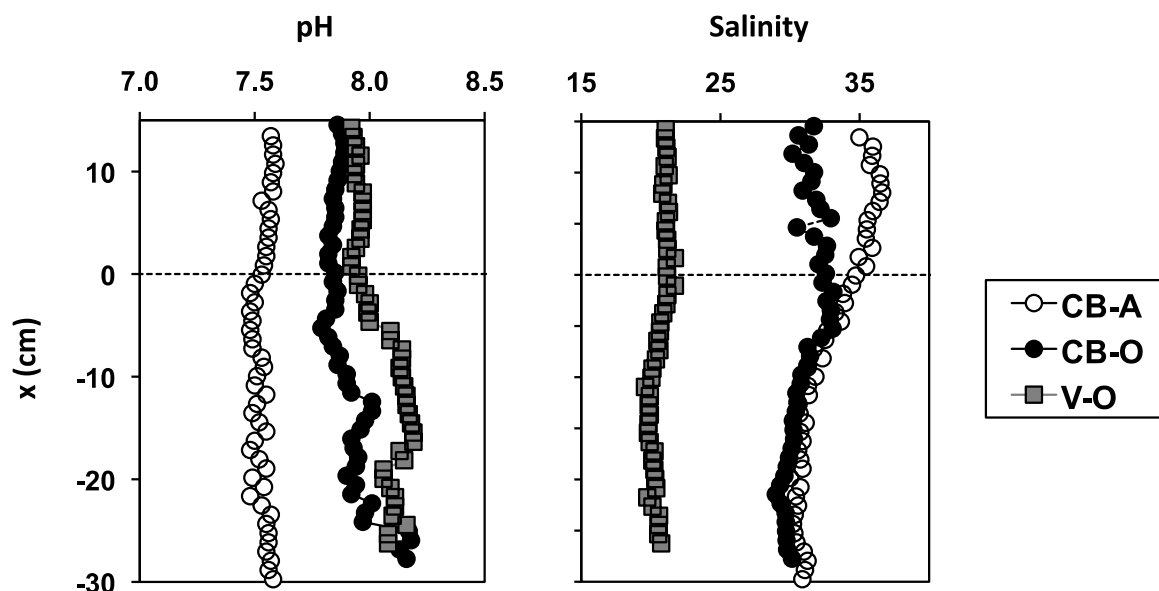


Figure EA-1.1: pH and salinity profiles in the overlying and pore waters collected by peepers. Salinity was obtained from measured Na concentrations and Dittmar's principle ($\text{Salinity} = [\text{Na}(\text{mmol/kg})] \times 0.07510$). The horizontal dashed line represents the sediment-water interface.

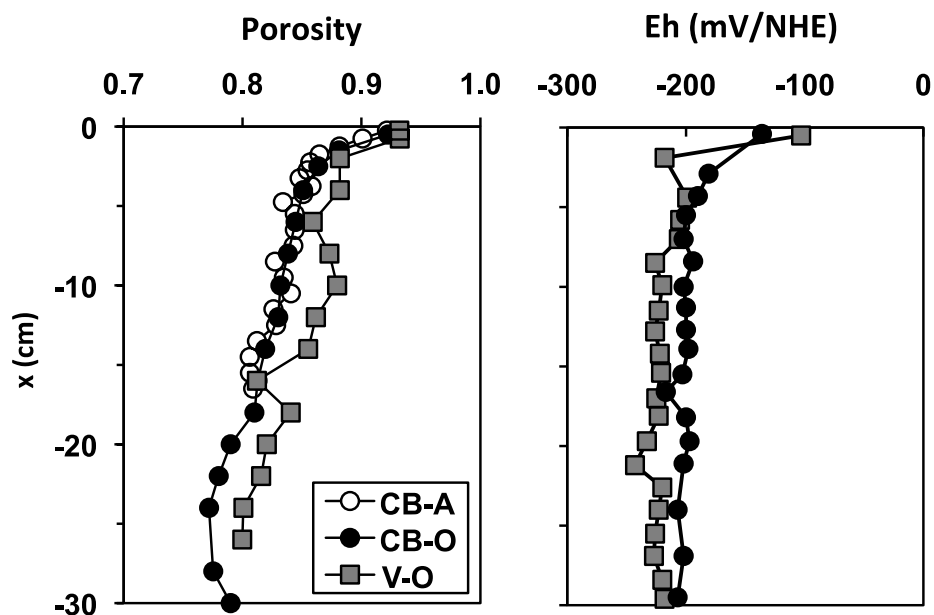
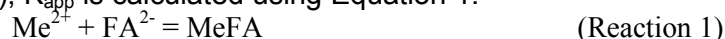


Figure EA-1.2: Porosity and Eh (expressed as mV normalized to the Standard Hydrogen Electrode) profiles in the sediment cores. The Eh profile was not carried out at anoxic site CB-A and was fixed at -200 mV for thermodynamic calculations.

EA-2: Calculation of apparent thermodynamic constants (K_{app}) of trace elements complexation with dissolved organic matter using WHAM-VI

The apparent thermodynamic constants (K_{app}) of trace elements complexation with fulvic acids (FA) and humic acids (HA) were calculated from the speciation calculated by WHAM VI. For example, for a complexation reaction between the metal Me^{2+} and the fulvic acid FA^{2-} forming a complex MeFA (Reaction 1), K_{app} is calculated using Equation 1.



$$\log K_{app} = \log \frac{[MeFA]}{[FA^{2-}][Me^{2+}]} \quad (\text{Equation 1})$$

where $[MeFA]$, $[FA^{2-}]$ and $[Me^{2+}]$ are the concentrations (mol.l^{-1}) of complex MeFA, fulvic acid FA^{2-} and free metal Me^{2+} in solution respectively. The molar concentrations of FA and HA in solution were calculated from their respective molecular weights, 650 and 2000 g.mol^{-1} , and the concentrations of dissolved organic carbon (DOC in g.l^{-1}) according to equations 2 and 3.

$$[FA^{2-}] = 2 \frac{DOC}{650} 0.9 \quad (\text{Equation 2})$$

$$[HA^{2-}] = 2 \frac{DOC}{2000} 0.1 \quad (\text{Equation 3})$$

The K_{app} were calculated for all samples in solution obtained in this work. Sulphide was shown to significantly modify the affinity of metals for dissolved organic matter, and two calculations were made: in the presence and absence of ΣH_2S . The K_{app} values used in this work correspond to the average of all measurements in both situation (with and without ΣH_2S) and are reported with associated standard deviation, in Table EA-1.1. The lack of data concerning the interaction between As and dissolved FA and HA, does not allow us to determine the K_{app} for this element.

Table EA-2.1: Averages and standard deviations of the apparent thermodynamic constants (K_{app}) reactions of trace elements complexation with fulvic acids (FA) and humic acids (HA) calculated from WHAM VI in the presence and absence of ΣH_2S . These reactions and K_{app} values were also added to the MINTEQ database.

Reactions	$\Sigma H_2S = 0 \mu M$		$\Sigma H_2S > 0 \mu M$	
	log K_{app}	s.d.	log K_{app}	s.d.
$Fe^{2+} + Fulvate^{2-} = FeFulvate$	3.5	0.2	3.6	0.1
$Fe^{2+} + Humate^{2-} = FeHumate$	4.3	0.1	4.5	0.2
$Mn^{2+} + Fulvate^{2-} = MnFulvate$	4.0	0.2	4.1	0.1
$Mn^{2+} + Humate^{2-} = MnHumate$	1.8	0.1	1.9	0.2
$Co^{2+} + Fulvate^{2-} = CoFulvate$	2.8	0.2	6.4	0.2
$Co^{2+} + Humate^{2-} = CoHumate$	3.5	0.1	6.5	0.2
$Cr^{3+} + Fulvate^{2-} = CrFulvate^+$	11.1	1.8	11.7	2.1
$Cr^{3+} + Humate^{2-} = CrHumate^+$	13.0	2.1	13.5	2.1
$Hg^{2+} + Fulvate^{2-} = HgFulvate$			32.7	0.3
$Hg^{2+} + Humate^{2-} = HgHumate$			35.8	1.3
$MeHg^+ + Fulvate^{2-} = MeHgFulvate^-$			10.3	0.2
$MeHg^+ + Humate^{2-} = MeHgHumate^-$			10.4	0.2
$Ni^{2+} + Fulvate^{2-} = NiFulvate$	3.4	0.2	5.6	0.2
$Ni^{2+} + Humate^{2-} = NiHumate$	3.9	0.1	5.7	0.2

EA-3: Reactions and associated thermodynamic constants (I = 0 M ; T = 25°C) added to the default WHAM-VI and MINTEQA2 databases for aqueous species speciation calculations.

N°	Reactions	Log k	References
<i>Carbonates</i>			
1	$\text{CO}_3^{2-} + \text{H}^+ = \text{HCO}_3^-$	9.37	Boudreau, 1997
2	$\text{CO}_3^{2-} + 2\text{H}^+ = \text{H}_2\text{CO}_3$	15.37	Boudreau, 1997
<i>Phosphates</i>			
3	$\text{PO}_4^{3-} + \text{H}^+ = \text{HPO}_4^{2-}$	12.34	Boudreau, 1997
4	$\text{PO}_4^{3-} + 2\text{H}^+ = \text{H}_2\text{PO}_4^-$	19.54	Boudreau, 1997
5	$\text{PO}_4^{3-} + 3\text{H}^+ = \text{H}_3\text{PO}_4$	21.69	Boudreau, 1997
<i>S(-2)/S(6)</i>			
6	$\text{HS}^- + \text{H}^+ = \text{H}_2\text{S}$	6.98	Boudreau, 1997
7	$\text{HS}^- = \text{S}^{2-} + \text{H}^+$	-18.57	Schoonen and Barnes, 1998 (Balistrieri, 1992)
8	$\text{HS}^- + 4\text{H}_2\text{O} = \text{SO}_4^{2-} + 9\text{H}^+ + 8\text{e}^-$	-34.0	Stumm and Morgan 1981 (Balistrieri, 1992)
<i>Fe</i>			
9	$\text{Fe}^{2+} + \text{HS}^- = \text{Fe}(\text{HS})^+$	5.2	Luther et al., 1996 (Rickard and Morse, 2005)
10	$\text{Fe}^{2+} + 2\text{HS}^- = \text{Fe}(\text{HS})_2$	8.542	Dyrssen, 1985 (Taillefert et al., 2000)
11	$\text{Fe}^{2+} + 3\text{HS}^- = \text{Fe}(\text{HS})_3^-$	9.808	Allison et al., 1989 (Taillefert et al., 2000)
12	$\text{Fe}^{2+} + 2\text{HS}^- = \text{FeHS}_2^- + \text{H}^+$	2.262	Dyrssen, 1985 (Taillefert et al., 2000)
<i>Mn</i>			
13	$\text{Mn}^{2+} + \text{HS}^- = \text{MnS} + \text{H}^+$	-2.08	Balistrieri, 1992
14	$\text{Mn}^{2+} + \text{HS}^- = \text{Mn}(\text{HS})^+$	-0.35	Balistrieri, 1992
15	$\text{Mn}^{2+} + 2\text{HS}^- = \text{Mn}(\text{HS})_2$	7.22	Balistrieri, 1992
16	$\text{Mn}^{2+} + 3\text{HS}^- = \text{Mn}(\text{HS})_3^-$	8.07	Balistrieri, 1992
17	$\text{Mn}^{2+} + 4\text{HS}^- = \text{Mn}(\text{HS})_4^{2-}$	10.61	Balistrieri, 1992
<i>As</i>			
18	$\text{AsO}_3^{3-} + \text{H}^+ = \text{HASO}_3^{2-}$	15	Nordstrom and Archer, 2003 (Couture et al., 2010)
19	$\text{AsO}_3^{3-} + 2\text{H}^+ = \text{H}_2\text{AsO}_3^-$	29.2	Nordstrom and Archer, 2003 (Couture et al., 2010)
20	$\text{AsO}_3^{3-} + 3\text{H}^+ = \text{H}_3\text{AsO}_3$	38.3	Nordstrom and Archer, 2003 (Couture et al., 2010)
21	$\text{AsO}_4^{3-} + \text{H}^+ = \text{HASO}_4^{2-}$	11.80	Nordstrom and Archer, 2003 (Couture et al., 2010)
22	$\text{AsO}_4^{3-} + 2\text{H}^+ = \text{H}_2\text{AsO}_4^-$	18.79	Nordstrom and Archer, 2003 (Couture et al., 2010)
23	$\text{AsO}_4^{3-} + 3\text{H}^+ = \text{H}_3\text{AsO}_4$	21.09	Nordstrom and Archer, 2003 (Couture et al., 2010)
24	$\text{AsO}_4^{3-} + 2\text{H}^+ + \text{Fe}^{2+} = \text{FeH}_2\text{AsO}_4^+$	5.15	Whiting, 1992 (Couture et al., 2010)
25	$\text{AsO}_4^{3-} + \text{H}^+ + \text{Fe}^{2+} = \text{FeHASO}_4$	6.1	Whiting, 1992 (Couture et al., 2010)
26	$\text{AsO}_4^{3-} + \text{Fe}^{2+} = \text{FeAsO}_4^-$	-13.6	Whiting, 1992 (Couture et al., 2010)
27	$\text{AsO}_4^{3-} + \text{H}^+ + \text{Mn}^{2+} = \text{MnHASO}_4$	15.54	Whiting, 1992 (Couture et al., 2010)
28	$\text{AsO}_4^{3-} + \text{Mn}^{2+} = \text{MnAsO}_4^-$	6.1	Whiting, 1992 (Couture et al., 2010)
29	$\text{AsO}_4^{3-} + \text{Mg}^{2+} = \text{MgAsO}_4^-$	6.3	Whiting, 1992 (Couture et al., 2010)
30	$\text{AsO}_4^{3-} + \text{H}^+ + \text{Mg}^{2+} = \text{MgHASO}_4$	14.7	Whiting, 1992 (Couture et al., 2010)
31	$\text{AsO}_4^{3-} + 2\text{H}^+ + \text{Ca}^{2+} = \text{CaH}_2\text{AsO}_4^+$	20.09	Bothe and Brown, 1999 (Couture et al., 2010)
32	$\text{AsO}_4^{3-} + \text{H}^+ + \text{Ca}^{2+} = \text{CaHASO}_4$	14.46	Bothe and Brown, 1999 (Couture et al., 2010)
33	$\text{AsO}_4^{3-} + \text{Ca}^{2+} = \text{CaAsO}_4^-$	4.36	Bothe and Brown, 1999 (Couture et al., 2010)
34	$\text{AsO}_3^{3-} + \text{HS}^- + 4\text{H}^+ = \text{H}_3\text{AsSO}_2 + \text{H}_2\text{O}$	49.88	Vlassopoulos et al., 2010
35	$\text{AsO}_3^{3-} + \text{HS}^- + 3\text{H}^+ = \text{H}_2\text{AsSO}_2^- + \text{H}_2\text{O}$	44.62	Vlassopoulos et al., 2010
36	$\text{AsO}_3^{3-} + 2\text{HS}^- + 3\text{H}^+ = \text{HAS}_2\text{O}_2^- + 2\text{H}_2\text{O}$	46.9	Wilkin et al., 2003 (Couture et al., 2010)
37	$\text{AsO}_3^{3-} + 3\text{HS}^- + 4\text{H}^+ = \text{HAS}_3\text{O}_2^- + 3\text{H}_2\text{O}$	58.7	Wilkin et al., 2003 (Couture et al., 2010)
38	$\text{AsO}_3^{3-} + 3\text{HS}^- + 3\text{H}^+ = \text{AsS}_3^{3-} + 3\text{H}_2\text{O}$	50.8	Wilkin et al., 2003 (Couture et al., 2010)
39	$\text{AsO}_3^{3-} + 4\text{HS}^- + 6\text{H}^+ = \text{As}(\text{SH})_4^- + 3\text{H}_2\text{O}$	74.8	Wilkin et al., 2003 (Couture et al., 2010)
40	$\text{AsO}_3^{3-} + 3\text{HS}^- + 6\text{H}^+ = \text{H}_3\text{AsS}_3 + 3\text{H}_2\text{O}$	66.91	Vlassopoulos et al., 2010
41	$\text{AsO}_3^{3-} + 3\text{HS}^- + 5\text{H}^+ = \text{H}_2\text{AsS}_3^- + 3\text{H}_2\text{O}$	64.77	Vlassopoulos et al., 2010
42	$\text{AsO}_3^{3-} + 3\text{HS}^- + 4\text{H}^+ = \text{HAS}_3\text{S}_2^- + 3\text{H}_2\text{O}$	57.58	Vlassopoulos et al., 2010
<i>Co</i>			
43	$\text{Co}^{2+} + \text{S}^{2-} = \text{CoS}$	21.88	Balistrieri, 1992
44	$\text{Co}^{2+} + \text{S}^{2-} + \text{H}^+ = \text{Co}(\text{HS})^+$	23.61	Balistrieri, 1992
45	$\text{Co}^{2+} + 2\text{S}^{2-} + 2\text{H}^+ = \text{Co}(\text{HS})_2$	49.75	Balistrieri, 1992
46	$\text{Co}^{2+} + 3\text{S}^{2-} + 3\text{H}^+ = \text{Co}(\text{HS})_3^-$	70.17	Balistrieri, 1992
47	$\text{Co}^{2+} + 4\text{S}^{2-} + 4\text{H}^+ = \text{Co}(\text{HS})_4^{2-}$	90.29	Balistrieri, 1992
<i>Hg</i>			
48	$\text{Hg}^{2+} + 2\text{HS}^- = \text{Hg}(\text{HS})_2$	40.37	Jay et al., 2000 (Feyte et al., 2011)
49	$\text{Hg}^{2+} + 2\text{HS}^- + \text{OH}^- = \text{HgS}(\text{HS})^- + \text{H}_2\text{O}$	48.6	Jay et al., 2000 (Feyte et al., 2011)
50	$\text{Hg}^{2+} + 2\text{HS}^- + 2\text{OH}^- = \text{HgS}_2^{2-} + 2\text{H}_2\text{O}$	53.56	Jay et al., 2000 (Feyte et al., 2011)
51	$\text{Hg}^{2+} + \text{HS}^- = \text{HgSH}^+$	22.3	Jay et al., 2000 (Feyte et al., 2011)

52	$\text{Hg}^{2+} + \text{HS}^- + \text{OH}^- = \text{HgS} + \text{H}_2\text{O}$	43.8	Jay et al., 2000 (Feyte et al., 2011)
53	$\text{Hg}^{2+} + \text{OH}^- = \text{HgOH}^+$	10.6	Powell et al., 2005 (Feyte et al., 2010)
54	$\text{Hg}^{2+} + 2\text{OH}^- = \text{Hg}(\text{OH})_2$	22.02	Powell et al., 2005 (Feyte et al., 2010)
55	$\text{Hg}^{2+} + 3\text{OH}^- = \text{Hg}(\text{OH})_3^-$	20.9	Powell et al., 2005 (Feyte et al., 2010)
56	$\text{Hg}^{2+} + \text{Cl}^- + \text{OH}^- = \text{HgClOH}$	18.27 ^a	Martell et al., 1998 (Merritt and Amirbahman, 2007)
57	$\text{Hg}^{2+} + \text{Cl}^- = \text{HgCl}^+$	7.3	Powell et al., 2005 (Feyte et al., 2010)
58	$\text{Hg}^{2+} + 2\text{Cl}^- = \text{HgCl}_2$	14.0	Powell et al., 2005 (Feyte et al., 2010)
59	$\text{Hg}^{2+} + 3\text{Cl}^- = \text{HgCl}_3^-$	14.93	Powell et al., 2005 (Feyte et al., 2010)
60	$\text{Hg}^{2+} + 4\text{Cl}^- = \text{HgCl}_4^{2-}$	15.5	Powell et al., 2005 (Feyte et al., 2010)
61	$\text{Hg}^{2+} + \text{SO}_4^{2-} = \text{HgSO}_4$	2.6	Powell et al., 2005 (Feyte et al., 2010)
62	$\text{Hg}^{2+} + \text{CO}_3^{2-} = \text{HgCO}_3$	11.51	Powell et al., 2005 (Feyte et al., 2010)
63	$\text{Hg}^{2+} + \text{OH}^- + \text{CO}_3^{2-} = \text{Hg}(\text{OH})\text{CO}_3^-$	19.34	Powell et al., 2005 (Feyte et al., 2010)
64	$\text{Hg}^{2+} + \text{H}^+ + \text{CO}_3^{2-} = \text{HgHCO}_3^+$	15.84	Powell et al., 2005 (Feyte et al., 2010)
65	$\text{Hg}^{2+} + 2\text{CO}_3^{2-} = \text{Hg}(\text{CO}_3)_2^{2-}$	15.58	Martell et al., 2001 (Feyte et al., 2011)
<i>MeHg</i>			
66	$\text{MeHg}^+ + \text{HS}^- = \text{MeHgHS}$	14.5	Loux et al., 2007 (Feyte et al., 2011)
67	$\text{MeHg}^+ + \text{HS}^- = \text{MeHgS}^- + \text{H}^+$	7.0 ^a	Martell et al., 2003 (Goulet et al., 2007)
68	$2\text{MeHg}^+ + \text{HS}^- = (\text{MeHg})_2\text{S} + \text{H}^+$	23.3 ^a	Martell et al., 2003 (Goulet et al., 2007)
69	$3\text{MeHg}^+ + \text{HS}^- = (\text{MeHg})_3\text{S} + \text{H}^+$	30.3 ^a	Martell et al., 2003 (Goulet et al., 2007)
70	$\text{MeHg}^+ + \text{OH}^- = \text{MeHgOH}$	9.47	De Robertis et al., 1998 (Feyte et al., 2010)
71	$2\text{MeHg}^+ + \text{OH}^- = (\text{MeHg})_2\text{OH}^+$	11.85	De Robertis et al., 1998 (Feyte et al., 2010)
72	$\text{MeHg}^+ + \text{CO}_3^{2-} = \text{MeHgCO}_3^-$	6.1	Rabenstein et al., 1976 (Feyte et al., 2010)
73	$\text{MeHg}^+ + \text{H}^+ + \text{CO}_3^{2-} = \text{MeHgHCO}_3$	12.95	Loux et al., 2007 (Feyte et al., 2010)
74	$\text{MeHg}^+ + \text{Cl}^- = \text{MeHgCl}$	5.45	De Robertis et al., 1998 (Feyte et al., 2010)
75	$\text{MeHg}^+ + \text{SO}_4^{2-} = \text{MeHgSO}_4^-$	2.64	De Robertis et al., 1998 (Feyte et al., 2010)
76	$\text{MeHg}^+ + \text{HPO}_4^{2-} = \text{MeHgHPO}_4^-$	5.48	Amirbahman et al., 2002 (Tipping et al., 2007)
<i>Ni</i>			
77	$\text{Ni}^{2+} + \text{S}^{2-} = \text{NiS}$	20.99	Balistrieri, 1992
78	$\text{Ni}^{2+} + \text{S}^{2-} + \text{H}^+ = \text{Ni}(\text{HS})^+$	22.73	Balistrieri, 1992
79	$\text{Ni}^{2+} + 2\text{S}^{2-} + 2\text{H}^+ = \text{Ni}(\text{HS})_2$	48.87	Balistrieri, 1992
80	$\text{Ni}^{2+} + 3\text{S}^{2-} + 3\text{H}^+ = \text{Ni}(\text{HS})_3^-$	69.29	Balistrieri, 1992
81	$\text{Ni}^{2+} + 4\text{S}^{2-} + 4\text{H}^+ = \text{Ni}(\text{HS})_4^{2-}$	89.41	Balistrieri, 1992

(a) I = 0.01 M

EA-4: Dissolution/precipitation reactions and associated thermodynamic constants (I = 0 M ; T = 25°C) added to the default MINTEQA2 databases for saturation indexes calculations.

N°	Minerals	Reactions	Log k	References
<i>C</i>				
1	Calcite	$\text{CaCO}_3 = \text{Ca}^{2+} + \text{CO}_3^{2-}$	-8.48	Boudreau, 1997
2	Aragonite	$\text{CaCO}_3 = \text{Ca}^{2+} + \text{CO}_3^{2-}$	-8.30	Boudreau, 1997
<i>S</i>				
3	Sulfur	$\text{S} + \text{H}^+ + 2\text{e}^- = \text{HS}^-$	-2.2	Jacob and Emerson, 1982 (Huerta Diaz et al., 1998)
<i>Fe</i>				
4	Siderite	$\text{FeCO}_3 = \text{Fe}^{2+} + \text{CO}_3^{2-}$	-10.8	Martell et al., 2001 (Gallon et al., 2004)
5	Ferrihydrite	$\text{Fe}(\text{OH})_3 + 3\text{H}^+ = \text{Fe}^{3+} + 3\text{H}_2\text{O}$	3.19	Parkhurst et Appelo, 1999
6	$\text{Fe}(\text{OH})_2$	$\text{Fe}(\text{OH})_2 = \text{Fe}^{2+} + 2\text{OH}^-$	-14.43	Martell et al., 2001 (Gallon et al., 2004)
7	FeS(am)	$\text{FeS} + \text{H}^+ = \text{Fe}^{2+} + \text{HS}^-$	-2.95	Davison, 1991 (Huerta Diaz et al., 1998)
8	Mackinawite	$\text{FeS} + \text{H}^+ = \text{Fe}^{2+} + \text{HS}^-$	-3.6	Davison, 1991 (Huerta Diaz et al., 1998)
9	Greigite	$\text{Fe}_3\text{S}_4 + 4\text{H}^+ = 2\text{Fe}^{3+} + \text{Fe}^{2+} + 4\text{HS}^-$	-34.69	Davison, 1991 (Huerta Diaz et al., 1998)
10	Pyrrhotite	$\text{FeS} + \text{H}^+ = \text{Fe}^{2+} + \text{HS}^-$	-5.1	Davison, 1991 (Huerta Diaz et al., 1998)
11	Troilite	$\text{FeS} + \text{H}^+ = \text{Fe}^{2+} + \text{HS}^-$	-5.25	Davison, 1991 (Huerta Diaz et al., 1998)
12	Pyrite	$\text{FeS}_2 + 2\text{H}^+ + 2\text{e}^- = \text{Fe}^{2+} + 2\text{HS}^-$	-11.47	Davison, 1991 (Huerta Diaz et al., 1998)
<i>Mn</i>				
13	Rhodochrosite	$\text{MnCO}_3 = \text{Mn}^{2+} + \text{CO}_3^{2-}$	-10.58	NIST, 1993 (Huerta Diaz et al., 1998)
14	MnS (Green)	$\text{MnS} + \text{H}^+ = \text{Mn}^{2+} + \text{HS}^-$	0.4	Smith and Martell, 1977 (Huerta Diaz et al., 1998)
15	MnS (pink)	$\text{MnS} + \text{H}^+ = \text{Mn}^{2+} + \text{HS}^-$	3.4	Smith and Martell, 1977 (Huerta Diaz et al., 1998)
16	Alabandite	$\text{MnS} + \text{H}^+ = \text{Mn}^{2+} + \text{HS}^-$	0.17	Dyrssen and Kremling, 1990 (Balistrieri, 1992)
17	Haurite	$\text{MnS}_2 + 2\text{H}^+ + 2\text{e}^- = \text{Mn}^{2+} + 2\text{HS}^-$	-3.7	Dyrssen and Kremling, 1990 (Balistrieri, 1992)
<i>As</i>				
18	As ₂ O ₅	$\text{As}_2\text{O}_5 + 3\text{H}_2\text{O} = 2\text{AsO}_4^{3-} + 6\text{H}^+$	-61.95	Nordstrom and Archer, 2003 (Couture et al., 2010)
19	Fe ₃ (AsO ₄) ₂	$\text{Fe}_3(\text{AsO}_4)_2 = 2\text{AsO}_4^{3-} + 3\text{Fe}^{2+}$	-34.41	Essington, 1988 (Couture et al., 2010)
20	Mn ₃ (AsO ₄) ₂ :8H ₂ O	$\text{Mn}_3(\text{AsO}_4)_2 \cdot 8\text{H}_2\text{O} = 2\text{AsO}_4^{3-} + 3\text{Mn}^{2+} + 8\text{H}_2\text{O}$	-29.02	Essington, 1988 (Couture et al., 2010)
21	Mg ₃ (AsO ₄) ₂	$\text{Mg}_3(\text{AsO}_4)_2 = 2\text{AsO}_4^{3-} + 3\text{Mg}^{2+}$	-30.02	Essington, 1988 (Couture et al., 2010)
22	Ca ₃ (AsO ₄) ₂ :4H ₂ O	$\text{Ca}_3(\text{AsO}_4)_2 \cdot 4\text{H}_2\text{O} = 2\text{AsO}_4^{3-} + 3\text{Ca}^{2+} + 4\text{H}_2\text{O}$	-21.15	Bothe and Brown, 1999 (Couture et al., 2010)
23	Ca ₅ (OH)(AsO ₄) ₃	$\text{Ca}_5(\text{OH})(\text{AsO}_4)_3 = 3\text{AsO}_4^{3-} + 5\text{Ca}^{2+} + \text{OH}^-$	-38.3	Bothe and Brown, 1999 (Couture et al., 2010)
24	CaHAsO ₄	$\text{CaHAsO}_4 = \text{AsO}_4^{3-} + \text{Ca}^{2+} + \text{H}^+$	-16.34	Bothe and Brown, 1999 (Couture et al., 2010)
25	Ferrarisite	$\text{Ca}_5(\text{HAsO}_4)_2(\text{AsO}_4)_2 = 4\text{AsO}_4^{3-} + 5\text{Ca}^{2+} + 2\text{H}^+$	-56.61	Bothe and Brown, 1999 (Couture et al., 2010)
26	Guerinite	$\text{Ca}_5(\text{HAsO}_4)_2(\text{AsO}_4)_2 = 4\text{AsO}_4^{3-} + 5\text{Ca}^{2+} + 2\text{H}^+$	-55.81	Bothe and Brown, 1999 (Couture et al., 2010)
27	Arsenolite	$\text{As}_2\text{O}_3 + 3\text{H}_2\text{O} = 2\text{AsO}_3^{3-} + 6\text{H}^+$	-77.92	Nordstrom and Archer, 2003 (Couture et al., 2010)
28	Claudetite	$\text{As}_2\text{O}_3 + 3\text{H}_2\text{O} = 2\text{AsO}_3^{3-} + 6\text{H}^+$	-77.95	Nordstrom and Archer, 2003 (Couture et al., 2010)
29	Orpiment	$\text{As}_2\text{S}_3 + 6\text{H}_2\text{O} = 2\text{AsO}_3^{3-} + 9\text{H}^+ + 3\text{HS}^-$	-122.8	Nordstrom and Archer, 2003 (Couture et al., 2010)
30	Orpiment(am)	$\text{As}_2\text{S}_3 + 6\text{H}_2\text{O} = 2\text{AsO}_3^{3-} + 9\text{H}^+ + 3\text{HS}^-$	-121.4	Nordstrom and Archer, 2003 (Couture et al., 2010)
31	Realgar	$\text{AsS} + 2.5\text{H}_2\text{O} + 0.25\text{O}_2 = \text{AsO}_3^{3-} + 4\text{H}^+ + \text{HS}^-$	-36.9	Nordstrom and Archer, 2003 (Couture et al., 2010)
<i>Co</i>				
32	CoS(alpha)	$\text{CoS} + \text{H}^+ = \text{Co}^{2+} + \text{HS}^-$	-7.4	Smith and Martell, 1977 (Huerta Diaz et al., 1998)
33	CoS(beta)	$\text{CoS} + \text{H}^+ = \text{Co}^{2+} + \text{HS}^-$	-11.7	Smith and Martell, 1977 (Huerta Diaz et al., 1998)
34	Covellite	$\text{CuS} + \text{H}^+ = \text{Cu}^{2+} + \text{HS}^-$	-22.27	Shea and Helz, 1989 (Huerta Diaz et al., 1998)
35	Chalcocite	$\text{Cu}_2\text{S} + \text{H}^+ = 2\text{Cu}^+ + \text{HS}^-$	-17.3	Smith and Martell, 1977 (Huerta Diaz et al., 1998)
<i>Hg</i>				
36	Cinnabar	$\text{HgS} + \text{H}^+ = \text{Hg}^{2+} + \text{HS}^-$	-39.1	Martell et al., 2003 (Goulet et al., 2007)
37	Metacinnabar	$\text{HgS} + \text{H}^+ = \text{Hg}^{2+} + \text{HS}^-$	-36.4 ^a	Martell et al. 1998 (Merritt and Amirbahman, 2007)
<i>Ni</i>				
38	Millerite	$\text{NiS} + \text{H}^+ = \text{Ni}^{2+} + \text{HS}^-$	-9.23	Smith and Martel, 1977 (Huerta Diaz et al., 1998)
39	NiS(alpha)	$\text{NiS} + \text{H}^+ = \text{Ni}^{2+} + \text{HS}^-$	-5.50	Smith and Martel, 1977 (Huerta Diaz et al., 1998)
40	NiS(beta)	$\text{NiS} + \text{H}^+ = \text{Ni}^{2+} + \text{HS}^-$	-11.00	Smith and Martel, 1977 (Huerta Diaz et al., 1998)
41	NiS(gamma)	$\text{NiS} + \text{H}^+ = \text{Ni}^{2+} + \text{HS}^-$	-12.70	Smith and Martel, 1977 (Huerta Diaz et al., 1998)

(a) I = 1 M

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